

Preparation of Acid Solutions

Stock Solution (M)	Dilution	Acid (mL)
Acetic Acid*		
CH ₃ COOH	5.5M	345 mL
FW: 60.05	3.5M	173
10.0%, 17.4 M	3.5M	56
sp. gr. 1.05	0.5 M	29
	0.1 M	5.8
Hydrochloric Acid*		
HCl	5.5M	700 mL
FW: 36.5	3.5M	280
10.0%, 12.1 M	3.5M	80
sp. gr. 1.05	0.5 M	40
	0.1 M	8.0
Nitric Acid*		
HNO ₃	5.5M	380 mL
FW: 63.01	3.5M	190
10.0%, 12.4 M	3.5M	60
sp. gr. 1.42	0.5 M	30
	0.1 M	6.0
Phosphoric Acid*		
H ₃ PO ₄	5.5M	400 mL
FW: 98.00	3.5M	200
10.0%, 12.0 M	3.5M	66
sp. gr. 1.70	0.5 M	33
	0.1 M	6.6
Sulfuric Acid*		
H ₂ SO ₄	5.5M	500 mL [†]
FW: 98.08	3.5M	250
10.0%, 12.0 M	3.5M	70
sp. gr. 1.84	0.5 M	35
	0.1 M	7.0

* Always add acid to water. The addition of acid to water is an exothermic reaction. Use high temperature (e.g., Pyrex®) glassware.

† Excessively exothermic; additional cooling must be used to keep the solution cool.

‡ Other amount of water required to prepare one liter of solution.

Preparing Sulfuric Acid Solution?

Always **ADD ACID TO WATER**! A great amount of heat is liberated when sulfuric acid is added to water. The temperature of the solution will rise rapidly. In fact, the temperature may rise so fast that the solution will boil and possibly spatter a strongly acidic solution. Consider containing your mixing vessel in a bucket of ice to control the solution temperature. Always add the acid to water very slowly while stirring continuously.


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