

1. Which of the following is a property of a system at equilibrium?

- A The rate of the forward reaction is zero. B The rate of the forward reaction is a minimum.
- C The rate of the forward reaction is equal to the rate of the reverse reaction. D The concentrations of reactants and products are equal.
- E The rate of the reverse reaction is a maximum.

2. Consider the reaction:



A $K_c = \frac{[\text{CH}_4][\text{H}_2\text{O}]^2}{[\text{CO}_2][\text{H}_2]^4}$

B $K_c = \frac{[\text{CH}_4][\text{H}_2\text{O}]^2}{[\text{CO}_2][\text{H}_2]^4}$

C none of the above

D $K_c = \frac{[\text{CH}_4][2\text{H}_2\text{O}]}{[\text{CO}_2][\text{H}_2]^4}$

E $K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CH}_4][\text{H}_2\text{O}]}$

3. In which system will decreasing the pressure increase the equilibrium concentration of the species in bold?



D none of the above



4. _____ What is the effect of increasing the temperature on the position of equilibrium and the value of the equilibrium constant for the reaction:

Position of equilibrium **Value of equilibrium constant**

A unchanged unchanged

B shifts to the right increases

Chemical Equilibrium Quiz

**Christopher R. Hren, John T.
Moore, Peter J. Mikulecky**

Chemical Equilibrium Quiz:

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