

1. Which of the following is a property of a system at equilibrium?

- ☐ A The rate of the forward reaction is zero.
- ☐ B The rate of of the forward reaction is a minimum
- ☐ C The rate of the forward reaction is equal to the rate of the reverse reaction.
- ☐ D The concentrations of reactants and products are equal.
- ☐ E The rate of the reverse reaction is a maximum.

2. Consider the reaction:



- ☐ A  $K_c = \frac{[\text{CH}_4][\text{H}_2\text{O}]^2}{[\text{CO}_2][\text{H}_2]^4}$
- ☐ B  $K_c = \frac{[\text{CH}_4][\text{H}_2\text{O}]^2}{[\text{CO}_2][\text{H}_2]}$
- ☐ C none of the above
- ☐ D  $K_c = \frac{[\text{CH}_4][2\text{H}_2\text{O}]}{[\text{CO}_2][4\text{H}_2]}$
- ☐ E  $K_c = \frac{[\text{CO}_2][\text{H}_2]}{[\text{CH}_4][\text{H}_2\text{O}]}$

3. In which system will decreasing the pressure increase the equilibrium concentration of the species in bold?

- ☐ A  $\text{NO}(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{N}_2\text{O}_3(\text{g})$
- ☐ B  $\text{N}_2\text{O}_3(\text{g}) + \text{NO}(\text{g}) \rightleftharpoons 3\text{NO}_2(\text{g})$
- ☐ C  $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
- ☐ D none of the above
- ☐ E  $\text{CH}_4(\text{g}) + 2\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}_2(\text{g}) + 4\text{H}_2(\text{g})$

4. What is the effect of increasing the temperature on the position of equilibrium and the value of the equilibrium constant for the reaction:

**Position of equilibrium    Value of equilibrium constant**

- ☐ A unchanged unchanged
- ☐ B shifts to the right increases

# Chemical Equilibrium Quiz

**Christopher R. Hren, John T.  
Moore, Peter J. Mikulecky**



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