



Part III: Balance the following chemical equations

20. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{AlCl}_3$
21. $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{FeO} \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{O}_3 + \underline{\hspace{1cm}} \text{Fe}$
22. $\underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{NaCl}$
23. $\underline{\hspace{1cm}} \text{P}_4\text{O}_{10} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{H}_3\text{PO}_4$
24. $\underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{CaSO}_4$
25. $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
26. $\underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
27. $\underline{\hspace{1cm}} \text{N}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{HNO}_3$
28. $\underline{\hspace{1cm}} \text{H}_3\text{AsO}_4 \rightarrow \underline{\hspace{1cm}} \text{As}_2\text{O}_5 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
29. $\underline{\hspace{1cm}} \text{Pb} + \underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{C}_2\text{H}_5\text{Cl} \rightarrow \underline{\hspace{1cm}} \text{Pb}(\text{C}_2\text{H}_5)_4 + \underline{\hspace{1cm}} \text{NaCl}$

Part IV: In each of the following problems you are given some potential reaction. Determine what type of reaction would occur, predict the products, write the complete chemical equation and balance it!

30. Aluminum powder burns in the presence of oxygen to form ...
31. Magnesium metal reacts with nitric acid to form ...
32. Aqueous barium chloride reacts with sodium hydroxide to form ...
33. Butane burns in the presence of oxygen to form ...
34. Hydrogen gas reacts with chlorine gas to form ...
35. Aluminum powder reacts with iron(III) oxide to form ...

Chemical Reactions Study Guide

Iris Stovall



Chemical Reactions Study Guide:

Study Guide to Organic Chemical Reactions Milin Kurup, 2020-01-06 This Organic Study Guide and workbook was created by the author in hopes to help students have an organized collection of all reactions studied in Organic 1 and Organic 2 With multiple reactions in different categories an organized study guide can help students focus on certain reactions as needed Many students in this class also have hopes of using this study guide as an MCAT review guide this organized set can be useful for students to review before class exams and standardized test With comprehensive notes reactions and tips this study guide will help everyone succeed This was created by a renowned student Milin Kurup at the University of Florida double majoring in Microbiology and Cognitive and Behavioral Neuroscience As a student of Professor Dr Laura Peterson UF Chemistry Organic Division this guide included all notes a series of 140 chemical reactions and key mechanisms necessary for a comprehensive understanding of the subject [Organic 1 and 2 Chemical Reactions Study Guide](#) Milin

Kurup, 2020-04-22 This Organic Study Guide and workbook was created by the author in hopes to help students have an organized collection of all reactions studied in Organic 1 and Organic 2 Many students in this class also have hopes of using this study guide as an MCAT review guide this organized set can be useful for students to review before class exams and standardized test This was created by a renowned student Milin Kurup at the University of Florida double majoring in Microbiology and Cognitive and Behavioral Neuroscience As a student of Professor Dr Laura Peterson UF Chemistry Organic Division this guide included all notes a series of 140 chemical reactions and key mechanisms necessary for a comprehensive understanding of the subject **Chemistry Equations And Answers (Speedy Study Guides)** Speedy

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