

Chemical Reactivity

Major Concepts

- 1.1 Metal
- 1.2 Non-Metals

Time allocation

Teaching periods	07
Assessment periods	02
Weightage	10%

Students Learning Outcomes

Students will be able to:

- Show how cation and anion are related to the terms metals and non-metals.
- Explain Alkali metals are not found in the free state in nature.
- Explain the differences in ionization energies of alkali and alkaline earth metals.
- Describe position of sodium metal in the periodic table its simple properties and uses.
- Position of calcium and magnesium in the periodic table, their simple properties and uses.
- Differentiate between soft and hard metals (iron and sodium)
- Describe the inertness of noble metals.
- Identify commercial value of silver, gold and platinum.
- Compile some important reactions of halogens.
- Name some elements that exist in nature in uncombined form.

Introduction

The different kinds of materials around us exist in variety of forms. Things like aeroplanes, trains, building frames, automobiles or even different machines and tools, are due to different properties of various metals. The non-metals exist as gases, liquids and soft or hard solids. They occupy upper right positions in the Periodic Table. Carbon, nitrogen, phosphorus, oxygen, sulphur, most of the halogens and the noble gases are non-metals. They show a variety of chemical reactivities. They form different ionic and covalent compounds, many of which are solids or gases.

8.1 METALS

Metals are the elements (except hydrogen) which are electropositive and form cations by losing electrons. Metals can be categorized.

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