

Study Guide-Chemistry Semester 1

Chapter 1

1. What are the 5 types of Chemistry? (page 8)
Organic Chemistry, Inorganic Chemistry, Biochemistry, Analytic Chemistry, Physical Chemistry.
2. What is the difference between pure and applied chemistry? (page 9)
3. Name 2 areas in which chemistry improved your world? (page 12-17)
4. How did Lavoisier transform chemistry? (page 21)
5. What are the Steps of the Scientific Method? (page 22)
6. What is a hypothesis? (page 22)
7. What is the difference between a scientific theory and a scientific law? (page 23)
8. What are 3 steps in solving numeric problems? (page 29)
9. What is the difference between a Manipulated Variable and a Responding Variable (page 22).

Chapter 2

1. What is the difference between a Heterogeneous Mixture and a Homogeneous Mixture?
2. What is the difference between a Chemical change and a chemical property?
3. What is the Law of Conservation of Mass?

Chapter 3 – Scientific Measurement

1. State the SI units for length, mass, temperature, and time
2. express numbers using scientific notation (for example, 3.29×10^6)
3. metric units – conversions
 - a. be able to convert from meters (m) to km, cm, mm, μm , nm (p.74)
 - i. convert 0.044 km to meters
 - b. be able to convert from liters (L) to mL, cm^3 , and μL (p.75)
 - i. convert 15 cm^3 in liters
 - c. be able to convert from kilograms (kg) to g, mg, and μg (p.76)
 - i. convert 750 dg in grams

Chapter 4 – Atomic Structure

1. Define **atom**.
2. State the 4 ideas of Dalton's atomic theory (p.102)
3. List the three types of subatomic particles (p.104)
4. What is the electrical charge of each subatomic particle?
5. Describe the structure of the atom according to the Rutherford model.
6. What distinguishes atoms of one element from atoms of a different element?
7. What is the **atomic number**? What is the **mass number**?
8. How do you calculate the number of neutrons in an atom (p.111)?



9. For the element $^{80}_{35}\text{Br}$ what is the atomic number, the mass number, and the number of neutrons?
10. What is an **isotope**?
11. Why do isotopes have different mass numbers?
12. What is the correct name for a row across the periodic table?
13. What is the correct name for a vertical group down the periodic table?

Chemistry Study Guide Answer

**Denise Guinn, Rebecca Brewer, Rachel
C. Lum**



Chemistry Study Guide Answer:

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