

# Basic Chemistry Study Guide 2

## 1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's  $\delta^+$  region is attracted to a  $\delta^-$  region in another molecule.
3. Form between water molecules (Where are the  $\delta^+$  and  $\delta^-$  regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life.
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

## 2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions.
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

## 3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic**. **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

# Chemistry Study Guide Completion

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Zafar, Moaz Siddiqui, Rachna Rani, Reetika Gulati, Sonal Chauhan, Maukta Gigras, 2021-11-25

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Physics and Measurement Kinematics Laws of Motion Work Energy and Power Rotational Motion Gravitation Properties of Solids Mechanical Properties of Fluids Thermal Properties of Matter Thermodynamics Kinetic Theory of Gases Simple Harmonic Motion Wave Motion Electrostatics Capacitance Current Electricity Magnetic Effects of Current Magnetism EM Induction and AC electromagnetic Waves Ray Optics Wave Optics Dual Nature of Matter and Radiation Atoms Nuclear Physics and Radioactivity Electronic Devices Communication Systems

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Hydrogen s p d f Block Elements Coordination Compounds Environmental Chemistry Purification of Organic Compounds Some Basic Principles of Organic Chemistry Hydrocarbons Organic Compounds Containing Halogens Alcohols Phenols and Ether Aldehyde Ketones and Carboxylic Acid Organic Compounds Containing Nitrogen Polymers Biomolecules

Chemistry in Everyday Life

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