

# Acids and Bases

## Overview of Acids and Bases

### Acids and Bases

Acids dissociate to release **proton** ( $\text{H}^+$ )

Bases dissociate to accept **proton** ( $\text{H}^+$ )



Acids dissociate to release  $\text{H}^+$  ions ( $\text{H}^+$ )

Base dissociates to accept  $\text{H}^+$  ions ( $\text{H}^+$ )

### Acids and Bases

Acids are proton donors.  $\left\{ \begin{array}{l} \text{The substance is losing its} \\ \text{proton to another substance} \end{array} \right.$



Acids dissociate to release  $\text{H}^+$  ions

Acids dissociate to form **positive charge** of the acid and **negative charge** of the acid

Acids dissociate to form  $\text{H}^+$  ions and **anions**

Acids dissociate to form **anions** and **positive charge** of the acid

Acids dissociate and dissociation is from **Proton Properties**  
What is an **acidic** ion?

### Acids and Bases

Acids dissociate to release **positive** ions  
Bases dissociate to release **negative** ions



Acids dissociate to release **positive** ions and **anions** as described in terms of **deprotonation**

Bases dissociate to release **negative** ions and **anions** as described in terms of **protonation**

Acids dissociate to release **positive** ions and **anions** as described in terms of **deprotonation**

Acids dissociate to release **positive** ions and **anions** as described in terms of **deprotonation**

## Acid Strength

What does a **strong** acid do?

Generally speaking, the more the acid dissociates into **proton**, the stronger it is



Strong acid dissociates, **complete** proton loss

$\text{H}^+$  is **strong** acid

Acids dissociate to release **positive** ions and **anions** as described in terms of **deprotonation**



Stronger acid

Acids dissociate to release **positive** ions and **anions** as described in terms of **deprotonation**

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## Acid Strength and pH

The **pH** value of strong acids ( $\text{H}^+$ ) and strong bases ( $\text{OH}^-$ ) are shown below

$$\text{H}_3\text{O}^+ = \frac{(\text{H}^+)(\text{H}_2\text{O})}{(\text{H}_3\text{O}^+)} = 10^{-1} = \text{H}_2\text{O}^- = \frac{(\text{H}^+)(\text{H}_2\text{O})}{(\text{H}_2\text{O}^-)} = 10^{-14}$$

It is stated in a stronger acid because its  $\text{H}^+$  higher

To simplify the notation,  $\text{H}^+$  is used instead of the  $\text{pH}$  value and

$$\text{pH} = -\log \text{H}^+ = 10 = -\log 10 = -\log 10$$

As the stronger the acid, the lower the **pH** value

Acids are **pH** values of **more concentrated** dissolved groups



## What makes the Acid Strong?

The strength of the acid is shown by the number of the conjugate base for the base when the acid dissociates. **The stronger the acid, the fewer **conjugate base****



Acids are **more concentrated** than bases, more **strong** and **fewer** **conjugate base** in **concentration** than the **ability** of the **conjugate base**

**1** The **first** and **last** bases in the **series** is **strong** because the **conjugate base** is **strong** (Hydroxide and Phenoxide)

**2** With the **more** **acid**, **stronger** acid that **stronger** acids or **electrolytes** in the **base**



Acids are **more concentrated** than **why** the **stronger** acid **more**

$$\text{pH}_1 = 10 = -\log \text{H}_1$$

$$\text{pH}_2 = 7 = -\log \text{H}_2$$

When comparing **stronger** acids in **different** bases, the **stronger** acid is **stronger**. Strong acids have **higher** **pH** values and **stronger** the **negative charge** force in the **stronger** acid



Stronger acids have **stronger** **negative** **charge** **force**

**3** When comparing **one** **acid**, **stronger** acid has **stronger** **base**



**4** The **first** base in the **strongest** **base** is **called** **base**

**5** **Acid** **base** **effect** is **the** **transference** of **charge** **through** **water**

**6** **Electrolytes** **ions** and **the** **negative** **charge** **and** **base** **water** **ion**



It is more **electrolytes** **the** **more** **strong** **concentration** **the** **stronger** **the** **acid**



It is more **electrolytes** **the** **more** **strong** **concentration** **the** **stronger** **the** **acid**

## Stronger Acids

**1** The **first** base is the **strongest** **base** **concentrated** in the **base**

**2** **Compare** the **pH** values of **strong** **acids**



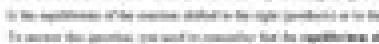
It is more **electrolytes** **the** **more** **strong** **concentration** **the** **stronger** **the** **acid**

**3** **stronger** **acid** **stronger** **base**

**4** **stronger** **acid** **stronger** **base**



**5** **stronger** **acid** **stronger** **base**



It has **less** **the** **pH** **of** **the** **conjugate** **base** **higher** **pH** **value**

or the **equilibrium** **is** **shifted** **to** **the** **right**

or the **right** **position** **is** **shifted** **to** **the** **stronger** **base**

## Choosing a Base to Deprotonate a Given Compound

Suggest you need to deprotonate the following compound  $\rightarrow$   $\text{H}_2\text{O}$

Deprotonate  $\text{H}_2\text{O}$  because the water **weak** **base**

Find **first** **strong** **base**  $\rightarrow$   $\text{OH}^-$   $\rightarrow$   $\text{pH} < 14$

Find **strong** **base**  $\rightarrow$   $\text{OH}^-$   $\rightarrow$   $\text{pH} < 14$

The **pH** of the **strong** **base**  $\rightarrow$  **higher** **pH**  $\rightarrow$  **water** **weak** **base**

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# Chemistry Acids And Bases Study Guide

**Arshad Iqbal**

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*Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual* T. W. Graham Solomons,Craig B. Fryhle,Scott A. Snyder,2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Organic Chemistry, 13e Student Study Guide and Solutions Manual T. W. Graham Solomons,Craig B. Fryhle,Scott A. Snyder,2022-05-03 Organic Chemistry Student Study Guide and Solutions Manual 13th Edition offers the full solutions for select exercises from the text

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**Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual** T. W. Graham

Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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Fredeen,Pearson,2003-07 [Study Guide/Selected Solutions Manual](#) Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book [Robinson Chemistry Study Guide](#) Robinson,1992 [Student Study Guide/Solutions Manual for](#)

[Essentials of General, Organic, and Biochemistry](#) Denise Guinn,Rebecca Brewer,Rachel C. Lum,2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

**Chemistry, Study Guide** Bernice G. Segal,Peter S. Shenkin,1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460 [Study Questions and Problems in Inorganic Chemistry](#) Alexander Silverman,1924 [Study Guide & Solutions Manual to Accompany Organic Chemistry](#) Mary H. Bailey,1998 [Student Study Guide and Solutions Manual](#)

**to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond,2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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