

Acids and Bases

Overview of Acids and Bases

Acid-base Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (proton acceptor)



Acid-base: (1) only in water, (2) only HBr is a base

What about organic molecules and inorganic substances?

Acid-base Acids and Bases

Acids are proton donors, Bases are proton acceptors.



1) Acid-base reactions of acid-base reactions:

to find the strong base, the base gets protonated, the acid gets protonated, the acid gets protonated, the acid gets protonated.

to find the strong base, the acid gets protonated, the acid gets protonated, the acid gets protonated, the acid gets protonated.

to find the strong base, the acid gets protonated, the acid gets protonated, the acid gets protonated, the acid gets protonated.

2) Acid-base and chemical definitions: are from Proton Properties. What if there is no proton transfer?

Acid-base Acids and Bases

Acids: acid is a proton donor, bases: base is a proton acceptor.



Acid-base: (1) only in water, (2) only HBr is a base

What about organic molecules and inorganic substances?

Acid Strength

What does it mean "strong acid"?

Acid-base: (1) only in water, (2) only HBr is a base



The dissociation of the acid is a reversible reaction that is described by the equilibrium constant.



Acid Strength and pKa

The Ka value of an acid (acid) and its conjugate base (B-) are related by:



The acid is a stronger acid than the base.

To compare the acid, the acid is a stronger acid than the base.

pKa = -logKa

$$\text{pKa} = -\log\text{Ka} = 14 - \text{pKb}$$

Acid-base: (1) only in water, (2) only HBr is a base

What about organic molecules and inorganic substances?



What makes the Acid Strong?

The strength of the acid is determined by the stability of the conjugate base. The more stable the conjugate base, the stronger the acid.

The stronger the acid, the more stable the conjugate base.



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1) The first factor for determining the acidity is called **inductive effect**.

Inductive effect is the transmission of charge through atoms.

Electronegativity atoms pull the negative charge and help to stabilize the conjugate base.



The more electronegative the atom, the stronger the acid.



The closer the electronegative atom to the hydrogen, the stronger the acid.

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Choosing a Base to Deprotonate a Given Compound

Suppose you want to deprotonate the following compound: —OH

Deprotonate means remove the proton (H+).

First, find the pKa value: —OH → pKa = 15

Find the chemical equation of the conjugate base with a conjugate base (B-).



The pKa of the conjugate acid must be higher than 15, so the acid is a strong acid.

Find a conjugate base with a pKa value of 15. Acids are weak when pKa > 15.



Choosing an Acid to Protonate a Given Compound

Suppose you want to protonate the following compound: —OH

Protonate means add the proton (H+).

To do that, find the chemical equation of the acid with a pKa value.



The pKa of the acid must be lower than 15, so the acid is a strong acid.

Find an acid with a pKa value of 15. Acids are weak when pKa > 15.

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Arshad Iqbal



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Table of Contents Chemistry Acids And Bases Study Guide

1. Understanding the eBook Chemistry Acids And Bases Study Guide
 - The Rise of Digital Reading Chemistry Acids And Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Acids And Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Acids And Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Acids And Bases Study Guide
 - Personalized Recommendations
 - Chemistry Acids And Bases Study Guide User Reviews and Ratings
 - Chemistry Acids And Bases Study Guide and Bestseller Lists
5. Accessing Chemistry Acids And Bases Study Guide Free and Paid eBooks
 - Chemistry Acids And Bases Study Guide Public Domain eBooks
 - Chemistry Acids And Bases Study Guide eBook Subscription Services
 - Chemistry Acids And Bases Study Guide Budget-Friendly Options

6. Navigating Chemistry Acids And Bases Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Acids And Bases Study Guide Compatibility with Devices
 - Chemistry Acids And Bases Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Acids And Bases Study Guide
 - Highlighting and Note-Taking Chemistry Acids And Bases Study Guide
 - Interactive Elements Chemistry Acids And Bases Study Guide
8. Staying Engaged with Chemistry Acids And Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Acids And Bases Study Guide
9. Balancing eBooks and Physical Books Chemistry Acids And Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Acids And Bases Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Acids And Bases Study Guide
 - Setting Reading Goals Chemistry Acids And Bases Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Acids And Bases Study Guide
 - Fact-Checking eBook Content of Chemistry Acids And Bases Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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