

ATOMIC STRUCTURE

RUTHERFORD'S MODEL

- Discovered that an atom is mostly empty space with tiny particles.
- Charged particles move in a small elliptical path around a central nucleus.
- This can only hold the negatively charged particles in the center.

BOHR'S MODEL

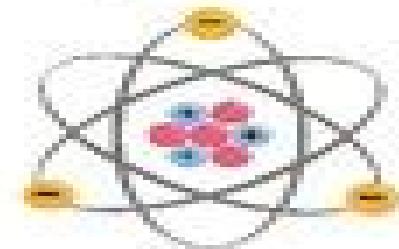
- Shows the elliptical path of an electron.
- Electrons
- Electrons

THE JELLY PUDDING MODEL



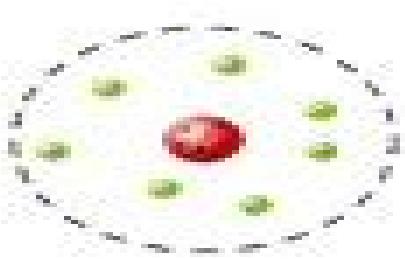
WAVE-MECHANICAL MODEL

- Shows electrons moving in paths, like tiny clouds like in clouds. Orbitals are regions of space likely where an electron.



WHAT IS TRUE TODAY...

- The nucleus is composed of 2 things:
 - PROTONS - positively charged particles
 - NEUTRONS - neutral charged particles
- Electrons are in a cloud - they are in orbitals
 - Electrons DO move around the nucleus in different patterns.



Subatomic Particle	Location in the Atom	Charge	Weight
proton	center (in the nucleus)	positive	1.67 x 10 ⁻²⁴
neutron	center (in the nucleus)	neutral (no charge)	1.67 x 10 ⁻²⁴
electron	around the nucleus in elliptical orbits	negative	1.67 x 10 ⁻²⁴ x 10 ⁻³⁰

IONS

- If an atom gains or loses an electron, it becomes an ion.
- This gives the atom a charge:
 - Loss of electrons, the atom becomes positive.
 - Gains of electrons, the atom becomes negative.

ISOTOPES

- An isotope is when an atom has a different mass number.
- An isotope is still the same element, it is just a little different mass, every atom has all other atoms different in slightly different mass.

Chemistry Atomic Structure Guide

Speedy Publishing

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A Guide to the Literature of Chemistry Evan Jay Crane, Austin McDowell Patterson, 1927

Chemistry

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Unveiling the Energy of Verbal Artistry: An Psychological Sojourn through **Chemistry Atomic Structure Guide**

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