

## General Chemistry II Chapter 17 Study Guide

### Organic acids- carboxylic acids ( $R\text{-COOH}$ ):

- Carboxylic acids ( $R\text{-COOH}$ ): a weak organic acid, contains the carb-oxyl group ( $-\text{COOH}$ ), never fully dissociates in water, an equilibrium process, ex. Formic acid, acetic acid, butyric acid, benzoic acid
  - $R$  = rest of the molecule
- Strength of carboxylic acids: depends on the  $R$  group; more electronegative  $\text{Cl}$  in  $R$  group shifts more electron density towards the  $R$  group and helps stabilize the structure after the  $\text{H}^+$  ion is removed, thus, chloroacetic acid is the stronger acid in comparison to acetic acid ( $\text{CH}_3\text{COOH}$ )

### Organic bases- amines:

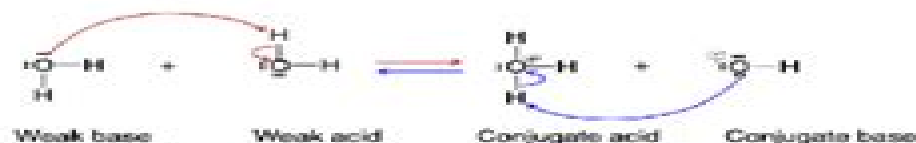
- Amines: organic bases, contains an  $\text{N}$  atom with a lone pair (ex.  $-\text{NH}$ ,  $-\text{NH}_2$ ), ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
  - Only  $\text{N}$  with lone pair can accept a proton
  - Equilibrium dissociation process
  - $\text{NH}_4^+$  is an acid and can lose one  $\text{H}^+$  to be neutral

### Organic acids and bases (things to keep in mind):

- Acid:  $\text{H}$  is the only acidic proton; ex.  $\text{CH}_3\text{COOH}$  becomes  $\text{CH}_3\text{COO}^-$
- Amine: the proton  $\text{H}^+$  is bonded to  $\text{N}$  via the lone pair of  $\text{N}$ ; ex.  $\text{CH}_3\text{NH}_2$  becomes  $\text{CH}_3\text{NH}_3^+$
- How about carbon?
  - Protons on  $\text{C}$  are not acidic;  $\text{C}$  cannot accept extra protons

### Autoionization of water:

- Autoionization: in pure water, one water molecule can donate a proton to another water molecule



$$\text{At } 25^\circ\text{C}, [\text{H}_3\text{O}^+] = [\text{OH}^-] = 1.0 \times 10^{-7} \text{ M}$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$K_w$  = ion-product constant

# Chemistry Chapter 17 Study Guide Answers

**Susan McMurry**



## Chemistry Chapter 17 Study Guide Answers:

**Study Guide and Solutions Manual** Neil E. Schore, 2002-08-02      **Organic Chemistry Study Guide and Solutions Manual** Marc Loudon, Jim Parise, Lisa Bonner, 2024-08-09 The Study Guide and Solutions Manual offers several study tools for students Study Guide Links are brief supplements to text material and provide hints for study shortcuts approaches to problem solving or additional explanations of topics that typically cause difficulty Further Explorations are short discussions that provide additional depth on content in the text Reaction Reviews present the reactions in the chapter in detailed outline form This summary consolidates the reactions of the chapter in one place and presents not only the reactions themselves but also the essential features of their mechanisms and stereochemistry for cases in which these issues are discussed in the text Solutions are provided for all of the over 1600 problems in the text Many include worked out examples in which not only the answer is given but also the detailed approach to the solution is discussed

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**Science Education** University of Illinois at Urbana-Champaign. Center for Instructional Research and Curriculum Evaluation,1978 **Study Guide and Solutions Manual for Fundamentals of Organic Chemistry** Susan McMurry,1994

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