

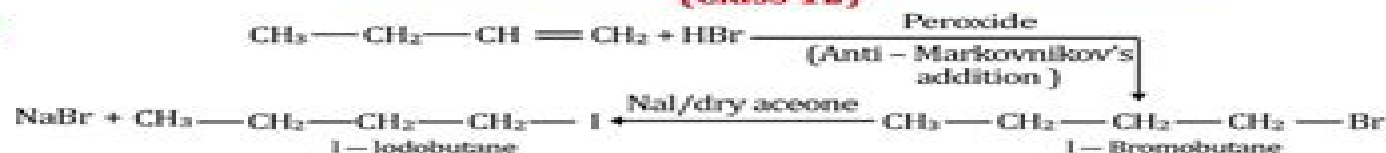
# Chemistry

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(Chapter 6) (Haloalkanes and Haloarenes) (Exercises)

(Class 12)

(iii)



**Question 6.8:**

What are ambident nucleophiles? Explain with an example.

**Answer 6.8:**

Ambident nucleophiles are nucleophiles having two nucleophilic sites. Thus, ambident nucleophiles have two sites through which they can attack. For example, nitrite ion is an ambident nucleophile.



Nitrite ion can attack through oxygen resulting in the formation of alkyl nitrites. Also, it can attack through nitrogen resulting in the formation of nitroalkanes.



**Question 6.9:**

Which compound in each of the following pairs will react faster in  $\text{S}_{\text{N}}2$  reaction with  $\text{OH}^-$ ?

(i)  $\text{CH}_3\text{Br}$  or  $\text{CH}_3\text{I}$

(ii)  $(\text{CH}_3)_3\text{CCl}$  or  $\text{CH}_3\text{Cl}$

**Answer 6.9:**

(i) In the  $\text{S}_{\text{N}}2$  mechanism, the reactivity of halides for the same alkyl group increases in the order. This happens because as the size increases, the halide ion becomes a better leaving group.



Therefore,  $\text{CH}_3\text{I}$  will react faster than  $\text{CH}_3\text{Br}$  in  $\text{S}_{\text{N}}2$  reactions with  $\text{OH}^-$ .

(ii)



The  $\text{S}_{\text{N}}2$  mechanism involves the attack of the nucleophile at the atom bearing the leaving group. But, in case of  $(\text{CH}_3)_3\text{CCl}$ , the attack of the nucleophile at the carbon atom is hindered because of the presence of bulky substituents on that carbon atom bearing the leaving group. On the other hand, there are no bulky substituents on the carbon atom bearing the leaving group in  $\text{CH}_3\text{Cl}$ . Hence,  $\text{CH}_3\text{Cl}$  reacts faster than  $(\text{CH}_3)_3\text{CCl}$  in  $\text{S}_{\text{N}}2$  reaction with  $\text{OH}^-$ .

**Question 6.10:**

Predict all the alkenes that would be formed by dehydrohalogenation of the following halides with sodium ethoxide in ethanol and identify the major alkene:

(i) 1-Bromo-1-methylcyclohexane

(ii) 2-Chloro-2-methylbutane

(iii) 2,2,3-Trimethyl-3-bromopentane.

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**James E. Brady**



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