

Chemistry 105 Equation Sheet

Constants

$$\begin{aligned}
 a_0 &= 0.5291772059 \text{ \AA} \\
 c &= 2.99792458 \times 10^8 \text{ m/s} \\
 e &= 1.6021765 \times 10^{-19} \text{ C} \\
 \epsilon_0 &= 8.8541877 \times 10^{-12} \text{ F/m} \\
 \hbar &= 6.626 \times 10^{-34} \text{ J}\cdot\text{s} \\
 k_B &= 1.38065 \times 10^{-23} \text{ J/K} \\
 k_e &= 9.987551 \times 10^9 \frac{\text{N}\cdot\text{m}^2}{\text{C}^2} \\
 m_e &= 9.109393 \times 10^{-31} \text{ kg} \\
 m_p &= 1.6749273 \times 10^{-27} \text{ kg} \\
 m_\pi &= 1.6726217 \times 10^{-27} \text{ kg} \\
 N_A &= 6.022142 \times 10^{23} \text{ mol}^{-1} \\
 R &= 8.31447 \frac{\text{J}}{\text{mol}\cdot\text{K}} \\
 &= 0.0820575 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} \\
 &= 1.987207 \frac{\text{cal}}{\text{mol}\cdot\text{K}} \\
 R_N &= 1.0974 \times 10^7 \text{ m}^2
 \end{aligned}$$

Conversions (* = exact)

Length

$$\begin{aligned}
 1 \text{ in} &= 2.54 \text{ cm}^* \\
 1 \text{ ft} &= 12 \text{ inches}^* \\
 1 \text{ yd} &= 3 \text{ ft}^* \\
 1 \text{ m} &= 39.37 \text{ inches}^* \\
 1 \text{ mile} &= 5280 \text{ feet}^* \\
 1 \text{ mile} &= 1.609 \text{ km}
 \end{aligned}$$

Mass

$$\begin{aligned}
 1 \text{ lb} &= 453.59237 \text{ g} \\
 &= 16 \text{ oz avdp.}^* \\
 1 \text{ US lb} &= 12 \text{ troy ounces}^* \\
 1 \text{ ton} &= 2000 \text{ lb}^* \\
 1 \text{ amu} &= 1.660538 \times 10^{-24} \text{ g}
 \end{aligned}$$

Volume

$$\begin{aligned}
 1 \text{ US gal} &= 3.7854 \text{ L} = 4 \text{ US qt}^* \\
 1 \text{ US qt} &= 32 \text{ US fl oz.}^* \\
 1 \text{ US fluid oz.} &= 29.5735 \text{ mL} \\
 1 \text{ cm}^3 &= 0.001 \text{ L}^* \\
 1 \text{ in}^3 &= 16.4 \text{ cm}^3
 \end{aligned}$$

* = exact ν = frequency v = velocity

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Pressure

$$\begin{aligned}
 1 \text{ atm} &= 760 \text{ torr}^* \\
 &= 101325 \text{ Pa}^* \\
 &= 406.733 \text{ in H}_2\text{O} \\
 &= 14.6959 \text{ lb/in}^2 \text{ (psi)} \\
 1 \text{ torr} &= 1 \text{ mm Hg (at } 0^\circ\text{C})^* \\
 1 \text{ Pa} &= 1 \text{ kg}\cdot\text{m}^{-1}\cdot\text{s}^{-2} \\
 1 \text{ bar} &= 100,000 \text{ Pa}^*
 \end{aligned}$$

Force

$$1 \text{ N} = 1 \text{ kg}\cdot\text{m}\cdot\text{s}^{-2}$$

Power

$$1 \text{ Watt} = 1 \text{ J/s}$$

Energy

$$\begin{aligned}
 1 \text{ J} &= 1 \text{ kg}\cdot\text{m}^2/\text{s}^2 = 0.2390 \text{ cal} \\
 1 \text{ J} &= 1 \text{ V}\cdot\text{C} = 0.009869 \text{ L}\cdot\text{atm} \\
 1 \text{ eV} &= 1.602 \times 10^{-19} \text{ J} \\
 1 \text{ L}\cdot\text{atm} &= 101.3 \text{ J}
 \end{aligned}$$

Formulas

$$\begin{aligned}
 \rho &= m/V \\
 T_F &= \frac{9F}{5C} T_C + 32 \text{ }^\circ\text{F} \\
 T_C &= \frac{5F}{9C} T_F + 273.15 \text{ K} \\
 c &= \lambda \nu \\
 E &= h\nu \\
 \lambda &= \frac{h}{mv} \\
 E_{\text{kinetic}} &= \frac{1}{2}mv^2 \\
 E &= mc^2 \\
 \Phi &= h\nu - KE_{\text{electron}}
 \end{aligned}$$

$$E_N = -Z^2 \frac{R_N hc}{n^2}$$

$$\Delta x \Delta(mv) \geq \frac{\hbar}{4\pi}$$

$$\frac{1}{\lambda} = R_N \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$V(r) = \frac{Q_1 Q_2}{4\pi \epsilon_0 r^2}$$

$$F(r) = k_e \frac{Q_1 Q_2}{r^2}$$

$$\mu = Qr$$

$$\Delta E = q + w$$

$$\Delta E_{\text{int}} = 0$$

$$\Delta S_{\text{int}} \geq 0$$

$$S = k_B \ln W$$

$$q = mC\Delta T$$

$$w = -P_{\text{ext}} \Delta V$$

$$\Delta H = \Delta E + P\Delta V$$

$$\Delta S = \frac{Q_{\text{rev}}}{T}$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^* + RT \ln Q$$

$$\Delta H_{\text{rxn}} = \Sigma n H_f^{\text{products}} - \Sigma m H_f^{\text{reactants}}$$

$$\Delta H_{\text{rxn}} \approx \Sigma n D_{\text{products}} - \Sigma m D_{\text{reactants}}$$

$$PV = nRT$$

$$P_{\text{tot}} = P_1 + P_2 + P_3 + \dots$$

$$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\text{Time}_B}{\text{Time}_A} = \sqrt{\frac{M_B}{M_A}}$$

$$KE_{\text{ave}} = \frac{3}{2}RT$$

$$v_{\text{rms}} = \sqrt{v^2} = \sqrt{\frac{3RT}{M}}$$

$$\text{pX} = -\log X$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

Chemistry Equation Sheet

Forman A. Williams

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Combustion Theory Forman A. Williams,2018-03-05 Combustion

Theory delves deeper into the science of combustion than most other texts and gives insight into combustions from a molecular and a continuum point of view The book presents derivations of the basic equations of combustion theory and contains appendices on the background of subjects of thermodynamics chemical kinetics fluid dynamics and transport processes Diffusion flames reactions in flows with negligible transport and the theory of pre mixed flames are treated as are detonation phenomena the combustion of solid propellents and ignition extinction and flammability phenomena

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