

Name KEY \_\_\_\_\_ Test Date \_\_\_\_\_ Period \_\_\_\_\_

## CP Chemistry – Final Exam Review

Show your work and round using significant figure rules for all the problems.

1. What is chemistry?  
• Chemistry is the study of matter and the changes it undergoes.
2. Write the following in scientific notation:
 

a. .005	$5 \times 10^{-3}$	c. 53099000	$5.3099 \times 10^7$
b. 5050	$5.05 \times 10^3$	d. 0.00053030	$5.3030 \times 10^{-4}$
3. Write the following in standard notation:
 

a. $1.5 \times 10^3$	<b>1500</b>	c. $4 \times 10^{-5}$	<b>0.00004</b>
b. $3.75 \times 10^{-2}$	<b>0.0375</b>	d. $9.12 \times 10^7$	<b>91200000</b>
4. How many significant figures are in the following:
 

a. 0.02	<b>1</b>	c. 501	<b>3</b>
b. 501.0	<b>4</b>	d. 0.020	<b>2</b>
5. Calculate and report in correct significant figures:
 

a. $52.568 \text{ cm} + 689.45 \text{ cm} + 55.5694 \text{ cm} =$	<b>797.59 cm</b>	
b. $8426.200 \text{ g} \div 254 \text{ mol} =$	<b>33.2 g/mol</b>	
6. You measure a Calcium cube to have 2.01 cm sides and a mass of 11.6 grams:
  - a. What is the density as you calculate it?  
$$v = s^3 = (2.01 \text{ cm})^3 = 8.12 \text{ cm}^3$$
  
$$d = \frac{m}{v} = \frac{11.6 \text{ g}}{8.12 \text{ cm}^3} = 1.43 \text{ g/cm}^3$$
  - b. If calcium's accepted density is 1.54 g/cm<sup>3</sup>, what is your percent error?  
$$\%_{\text{error}} = \frac{|1.54 \text{ g/cm}^3 - 1.43 \text{ g/cm}^3|}{1.54 \text{ g/cm}^3} \times 100 = 7.14 \%$$
7. Perform the following conversions:
  - a. 53.1 mm to m  
$$53.1 \text{ mm} \times \frac{1 \text{ m}}{10^3 \text{ mm}} = 0.0531 \text{ m}$$
  - b. 18.4 kg to dag  
$$18.4 \text{ kg} \times \frac{1 \times 10^4 \text{ dag}}{1 \times 10^3 \text{ kg}} = 1840 \text{ dag}$$
  - c. 93.6 mi to km (1 mi = 1.6 km)  
$$93.6 \text{ mi} \times \frac{1.6 \text{ km}}{1 \text{ mi}} = 150. \text{ km}$$

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