

Chemistry Formulas

Dilution	$M_1V_1 = M_2V_2$	$pK_a \& pK_b$	$pK_a = -\log[K_a]$
Energy of a photon	$E_{\text{photon}} = hf = hc/\lambda$	$pK_a \& pK_b$	$pK_b = -\log[K_b]$
Pressure	$P = F/A$	$pK_a \& pK_b$	$pK_a + pK_b = 14$
Average Kinetic Energy	$KE_{\text{avg}} = \frac{3}{2}RT$	Neutralization Reaction	$n_A M_A V_A = n_B M_B V_B$
Ideal Gas Law	$PV = nRT$	Buffers	$pH = pK_a + \log [A^-] [HA]$
Combined Gas Law	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$	Enthalpy of Formation	$\Delta H_f = \Sigma n\Delta H_f^{\circ}(\text{product}) - \Sigma n\Delta H_f^{\circ}(\text{reactants})$
Dalton's Law of Partial Pressures	$P_{\text{total}} = P_A + P_B + \dots$	First Law of Thermodynamics	$dE = q + w$
Dalton's Law of Partial Pressures	$P_A = X_A P_{\text{total}}$	Pressure Volume Work	$w = -P\Delta V$
Graham's Law of Effusion	$\frac{r_{\text{A}}}{r_{\text{B}}} = \sqrt{\frac{M_{\text{B}}}{M_{\text{A}}}}$	Calorimetry Thermal Energy (q)	$q = -C_{\text{calorimeter}} \Delta T$
Henry's Law	$P_A = k_H[A]$	Heat Curves & Thermal Energy (q)	$q = mc\Delta T$
Freezing Point Depression	$\Delta T_f = -iK_f M$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_f^{\circ}$
Boiling Point Elevation	$\Delta T_b = iK_b m$	Heat Curves & Thermal Energy (q)	$q = m\Delta H_{\text{vaporization}}$
Vapor Pressure Depression (Raoult's Law)	$P_{\text{total}} = X_{\text{water}} P^{\circ}_{\text{water}}$	Entropy (S)	$\Delta S = \Sigma nS_{\text{products}} - \Sigma nS_{\text{reactants}}$
Osmotic Pressure (π)	$\pi = iMRT$	Bond Dissociation Energy	$\Delta H = \Sigma n\Delta H_{\text{reactants}} - \Sigma n\Delta H_{\text{products}}$
Arrhenius Equation	$k = Ae^{-\nu\epsilon/RT}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$
Equilibrium Constant Expressions	$K_c = \frac{[\text{products}]}{[\text{reactants}]}$	Gibb's Free Energy (ΔG)	$\Delta G = \Delta G^{\circ} + RT\ln Q$
Equilibrium Constant Expressions	$K_{\text{eq}} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$	Gibb's Free Energy (ΔG)	$\Delta G^{\circ} = -RT^2 \ln(K_{\text{eq}})$
Equilibrium Constant Expressions	$K_P = \frac{P_{\text{products}}}{P_{\text{reactants}}}$	Standard Cell Potential	$E^{\circ} = E^{\circ}\text{reduction} + E^{\circ}\text{oxidation}$
Reaction Quotient (Q)	$Q = \frac{[\text{products}]}{[\text{reactants}]}$	Standard Cell Potential	$E^{\circ} = E^{\circ}\text{cathode} + E^{\circ}\text{anode}$
Solubility Product Constant (K_{sp})	$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$	Nernst Equation	$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$
pH & pOH	$pH = -\log[H^+]$	Kinetics (always 1st order)	$N = N_0 e^{-kt}$
pH & pOH	$pOH = -\log[OH^-]$	Kinetics (always 1st order)	$\ln N = \ln N_0 - kt$
pH & pOH	$pH + pOH = 14$	Kinetics (always 1st order)	$t_{1/2} = \frac{0.693}{k}$
		Nuclear Binding Energy	$E = \Delta mc^2$

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John Olmsted, Gregory M. Williams

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Chemical Formulas and Names

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U. Klein,2013-04-17 constitutive of reference in laboratory sciences as cultural sign systems and their manipulation and superposition collectively shared classifications and associated conceptual frameworks and various forms of collective action and social institutions This raises the question of how much modes of representation and specific types of sign systems mobilized to construct them contribute to reference Semioticians have argued that sign systems are not merely passive media for expressing preconceived ideas but actively contribute to meaning Sign systems are culturally loaded with meaning stemming from previous practical applications and social traditions of applications In new local contexts of application they not only transfer stabilized meaning but also can be used as active resources to add new significance and modify previous meaning This view is supported by several analyses presented in this volume Sign systems can be implemented like tools that are manipulated and superposed with other types of signs to forge new representations The mode of representation made possible by applying and manipulating specific types of representational tools such as diagrammatic rather than mathematical representations or Berzelian formulas rather than verbal language contributes to meaning and forges fine grained differentiations between scientists concepts Taken together the essays contained in this volume give us a multifaceted picture of the broad variety of modes of representation in nineteenth century and twentieth century laboratory sciences of the way scientists juxtaposed and

integrated various representations and of their pragmatic use as tools in scientific and industrial practice **Chemistry Equations and Answers** Mark Jackson,2006-02-22 This 6 page study guide contains basic chemistry analysis and concepts designed specifically to aid science students **Chemistry Equations And Answers (Speedy Study Guides)** Speedy Publishing,2015-04-24 Chemistry is a difficult subject to fully comprehend with its equations and scientific laws Trying to digest an entire book in one semester is a tough job but with the help of study guides like these you can absorb information in chemistry much more effectively This guide covers chemical equations including examples potential problems and solutions **Chemical Reactions and Their Equations** Ingo Waldemar Dagobert Hackh,1921 **Chemical Formulas** Alpha & Omega Publishing,2001-03-01 **Principles of chemistry** Joel Henry Hildebrand,1918 **Chemistry Equations & Answers** Speedy Publishing,2014-08-01 One of the most important parts of learning chemistry is simply knowing all of the equations and formulas that are used in it It is important to know the exact formatting of these equations and most classes will require a student to know them for exams It is a good idea to learn these equations and formulas with the use of a study pamphlet The pamphlet can condense all of the information so a student can memorize the equations and formulas while studying **Chemistry** Jason Overby,Raymond Chang,2024 The fifteenth edition continues a long tradition of providing a firm foundation in the concepts of chemical principles while instilling an appreciation of the important role chemistry plays in our daily lives We believe that it is our responsibility to assist both instructors and students in their pursuit of this goal by presenting a broad range of chemical topics in a logical format At all times we strive to balance theory and application and to illustrate principles with applicable examples whenever possible **An Introduction to the Principles of Physical Chemistry from the Standpoint of Modern Atomistics and Thermo-dynamics** Edward Wight Washburn,1915

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Chemical Structure, Spatial Arrangement Peter J. Ramberg,2017-05-15 Offering a comprehensive narrative of the early history of stereochemistry Dr Ramberg explores the reasons for and the consequences of the fundamental change in the meaning of chemical formulas with the emergence of stereochemistry during the last quarter of the nineteenth century As

yet relatively unexplored by historians the development of stereochemistry the study of the three dimensional properties of molecules provides a superb case study for exploring the meaning and purpose of chemical formulas as it entailed a significant change in the meaning of chemical formulas from the purely chemical conception of structure to the physico chemical conception of molecules provided by the tetrahedral carbon atom This study is the first to treat the emergence of the unique visual language of organic chemistry between 1830 and 1874 to place in context the near simultaneous proposal of the tetrahedral carbon atom by J H van t Hoff and J A Le Bel in 1874 Dr Ramberg then examines the research programs in stereochemistry by Johannes Wislicenus Arthur Hantzsch Victor Meyer Carl Bischoff Emil Fischer and Alfred Werner showing how the emergence of stereochemistry was a logical continuation of established research traditions in chemistry In so doing he also illustrates the novel and controversial characteristics of stereochemical ideas especially the unprecedented use of mechanistic and dynamic principles in chemical explanation

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