

Nuclear Chemistry - Guided Practice Questions**Multiple Choice** - Choose the one best answer and write its letter in the blank provided.

_____ 1) In balancing the nuclear reaction: $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + \text{_____}$, the identity of the missing element is _____?

- A) Pu B) U C) Th D) Np E) Pa

_____ 2) The following rxn, $^{210}_{84}\text{Po} \rightarrow ^{210}_{82}\text{Pb} + \text{_____}$, is an example of what form of radioactive decay?

- A) alpha B) positron C) gamma D) beta E) e-capture

_____ 3) The missing product from this reaction is: $^{138}_{54}\text{Xe} \rightarrow ^{138}_{52}\text{Te} + \text{_____}$?

- A) $^0_{-1}\text{e}$ B) $^0_0\gamma$ C) ^4_2He D) ^1_0n E) ^1_1p

_____ 4) What is the mass number of a neutron _____?

- A) 1 B) 3 C) 2 D) 0 E) 4

_____ 5) In the nuclear decay reaction represented by $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th}$, the emitted particle is _____.

- A) beta particle B) neutron C) alpha particle D) positron E) proton

_____ 6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?



- A) electron B) proton C) positron D) alpha particle E) neutron

_____ 7) $^{131}_{53}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 1.53 mg sample of $^{131}_{53}\text{I}$, how many mg will remain after 13.0 days _____?

- A) 0.440 B) 0.422 C) 0.268 D) 0.835 E) 0.499

_____ 8) All atoms of a given element have the same _____.

- A) atomic mass B) mass number C) # of nucleons D) atomic # E) # of neutrons

_____ 9) Atoms containing radioactive nuclei are called _____.

- A) radioisotopes B) radionuclides C) nuclides D) nucleons E) radionophores

_____ 10) What happens to the mass number and the atomic number of an element when it undergoes beta decay?

- A) The mass number decreases by 4 and the atomic number decreases by 2.
 B) The mass number does not change and the atomic number decreases by 2.
 C) Neither the mass number nor the atomic number change.
 D) The mass number increases by 2 and the atomic number increases by 1.
 E) The mass number does not change and the atomic number increases by 1.

Chemistry Guided Practice Problems Answer Key

John S. Phillips, Cheryl Wistrom



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Organic Chemistry, a Guided Inquiry Andrei Straumanis, 2004 Learning to Solve Complex Scientific Problems David H. Jonassen, 2017-09-25 Problem solving is implicit in the very nature of all science and virtually all scientists are hired retained and rewarded for solving problems Although the need for skilled problem solvers has never been greater there is a growing disconnect between the need for problem solvers and the educational capacity to prepare them Learning to Solve Complex Scientific Problems is an immensely useful read offering the insights of cognitive scientists engineers and science educators who explain methods for helping students solve the complexities of everyday scientific problems Important features of this volume include discussions on how problems are represented by the problem solvers and how perception attention memory and various forms of reasoning impact the management of information and the search for solutions how academics have applied lessons from cognitive science to better prepare students to solve complex scientific problems gender issues in science and engineering classrooms and questions to guide future problem solving research The innovative methods explored in this practical volume will be of significant value to science and engineering educators and researchers as well as to instructional designers **Invitation to Organic Chemistry** Alyn William Johnson, 1999 Colorful graphics and 19 chapters featuring such learning aids as chemistry at work and conceptual problems characterize this large text on a large subject Cited by the American Association for the Advancement of Science for his pioneering work in the chemistry of ylides Johnson who spent most of his career at the U of North Dakota explores the smorgasbord of subject matter that is organic chemistry and new developments in the field Appends a summary of nomenclature spectra group assignments and values of selected important compounds The index is combined with a glossary Annotation copyrighted by Book News Inc Portland OR **Keys to the Secondary Classroom** Rain Bongolan, Ellen Moir, Wendy Baron, 2009-11-16 The suggestions and tips give teachers a solid start to their year and help their students achieve This book has a unique focus on secondary teachers Janet Crews Secondary Instructional Coordinator Clayton School District MO A blueprint for successful teaching an entire year of teaching methods in one volume This book is jam packed with practical useful forms to guarantee a meaningful year for both the students and especially the teacher Ronald W Poplau Social Studies Teacher Shawnee Mission Northwest High School KS An easy to use source for all the strategies you need to thrive in the secondary classroom Preparing adolescent students for today's increasing demands requires commitment organization subject matter knowledge and passion But with the right strategies and tools first year secondary school teachers can approach this challenge with greater confidence and ease Leveraging a wealth of information and research from the New Teacher Center this user friendly guide provides a solid foundation for classroom management lesson planning and assessment Teachers will learn step by step tips for organizing standards based curriculum across the content areas supported by extensive reproducible forms and go to references including lesson plans from exemplary math and language arts teachers The authors provide specific strategies

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