

AP Chemistry Midterm Study Packet

Key Terms:

Chapter 14 – Liquids and Solids

- amorphous solid** — solid in which the arrangement of the representative particles lacks a regular, repeating pattern; do not have a definite melting point
- boiling point** — temperature at which the vapor pressure of a liquid becomes equal to atmospheric pressure
- condensation** — change of state from liquid to gas
- condensed state** — liquid and solid states of matter; states in which a substance has a substantially higher density than in the gaseous state
- covalent network solid** — solid in which strong covalent bonds form a network extending throughout the solid
- crystalline solid** — solid in which the representative particles are in a highly ordered, repeating pattern called a crystal
- deposition** — conversion of a gas directly into a solid without first becoming a liquid
- dipole-dipole force** — intermolecular force of attraction between neighboring permanent dipoles that is created when the dipoles line up so that the positive and negative ends are close to each other
- dispersion force** — intermolecular force of attraction between induced dipoles
- equilibrium vapor pressure** — pressure exerted by a vapor in equilibrium with its liquid; point at which the number of molecules in the vapor state remains solid
- evaporation** — process by which molecules of a liquid escape from the surface of the liquid and enter the gaseous, or vapor, state
- freezing point** — temperature at which the solid and liquid form of a substance exist in equilibrium
- heat of fusion** — heat necessary to convert a given amount of a solid into a liquid
- heat of vaporization** — amount of heat necessary to vaporize a given amount of liquid
- heating curve** — plot of the temperature of a sample as a function of time
- hydrogen bond** — strong intermolecular force between the hydrogen atom of one molecule and a highly electronegative atom of another molecule
- induced dipole** — dipole created by the presence of a neighboring dipole
- intermolecular force** — relatively weak force of attraction that exists between neighboring molecules
- intramolecular force** — force of attraction that exists within a molecule to hold it together
- phase diagram** — plot of the state of a sample as a function of temperature and pressure
- sublimation** — conversion of a solid directly into a gas, without first becoming a liquid
- surface tension** — imbalance of attractive forces at the surface of a liquid that causes the surface to behave as if it had a thin film across it; resistance of a liquid to an increase in the surface area
- vaporization** — change of state from liquid to gas
- viscosity** — resistance to motion that exists between the molecules of a liquid when they move past each other; affected by the extent of the hydrogen bonding

Chapter 15 – Solutions

- alloy** — solid solution in which the atoms of two or more metals are uniformly mixed
- aqueous solution** — solution in which the solvent is water
- boiling point elevation** — colligative property in which the boiling point of a solvent is raised when a nonvolatile solute is dissolved in the solvent; directly related to the concentration of the solution
- colligative property** — property of a solution that depends upon the concentration of the solute, but not its identity
- concentration** — amount of solute dissolved in a given amount of solvent
- freezing point depression** — colligative property in which the freezing point of a solvent is lowered when a nonvolatile solute is dissolved in the solvent; freezing point depression is directly related to the concentration of a solution
- hydration** — process in which water molecules pull solute particles into solution and form a sphere around them
- immiscible** — inability of a liquid to form a solution with another liquid in all proportions
- insoluble** — inability of a substance to dissolve in another substance
- miscible** — ability of a liquid to form a solution with another liquid in all proportions
- molarity** — concentration of a solution determined by the number of moles of solute per kilogram of solvent
- molarity** — concentration of a solution determined by the number of moles of solute per liter of solution
- mole fraction** — concentration of a solution determined by the number of moles of a given solute divided by the total number of moles of solution
- osmosis** — flow of solvent particles from a dilute solution to a concentrated solution
- osmotic pressure** — pressure required to prevent osmosis; arises when the flow of solvent particles from a dilute solution to a concentrated solution across a semipermeable membrane results in uneven heights of the solutions of either side of the membrane
- saturated** — description of a solution that contains as much dissolved solute as it can under existing conditions
- solubility** — amount of solute that dissolves in a given amount of solvent at a given temperature to form a saturated solution
- soluble** — ability of a solution to dissolve in another substance
- solute** — substance that is dissolved in a solvent to form a solution
- solution** — homogeneous mixture of two or more substances in a single physical state
- solvation** — process in which solvent particles pull solute particles into solution and form a sphere around them; called hydration if the solvent is water
- solvent** — substance that does the dissolving in a solution

Chemistry Midterm Study Guide

Cathy Duffy



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Student Study Guide Peter L. Kresan, Reed Mencke, Frank Press, Raymond Siever, 2003-09-25 This reconceptualization of the text *Understanding Earth* reflects the fundamental changes in the field of physical geology over the past several years

PSI Study Guide for Organic Chemistry by Douglas C. Neckers, Michael P. Doyle E. C. Blossey, 1977 *McGraw-Hill's 500 College Chemistry Questions* David E. Goldberg, 2012-09-07 500 Ways to Achieve Your Best Grades We want you to succeed on your college chemistry midterm and final exams That's why we've selected these 500 questions to help you study more effectively use your preparation time wisely and get your best grades These questions are similar to the ones you'll find on a typical college exam so you will know what to expect on test day Each question includes comprehensive explanations in the answer key Whether you have been studying all year or are doing a last minute review McGraw Hill's 500 College Chemistry Questions will help you achieve the final grade you desire Sharpen your subject knowledge and build your test taking confidence with 500 essential college chemistry questions with answers Clear solutions in the answer key for every problem Coverage from atomic mass to electrochemistry

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Making Pre-Med Count Elisabeth Fassas, 2020-01-07 In *Making Pre Med Count* med student Elisabeth Fassas shares

personal stories from her own experiences to help guide you through the pre med process You ll get first hand guidance and learn how to apply her advice to your own med school journey Counselors and checklists are helpful but your pre med journey cannot be boiled down to a list of activities and a collection of accolades In Making Pre Med Count Fassas teaches you how to translate your accomplishments into a compelling and personalized med school application Fassas gets into the weeds of the pre med years to touch on the most fundamental and gnawing questions that interested applicants must face Using examples from her own journey from freshman year to acceptance plus tips and tricks from her peers she guides readers through an endless stream of conflicting advice towards preparing academically mentally and psychologically for the med school application process Her advice starts with the idea that anything and everything can get you into medical school if you re able to get into the heads of the admissions committee You ll also get her take on many of the questions raised in student forums Fassas who will begin med school in the fall of 2019 helps relieve some of the common pre med doubts anxieties and fears that you ll feel Making Pre Med Count compiles Fassas advice in one place it s like having your own personal med school advisor

Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

Languaging Myths and Realities Qianqian Zhang-Wu, 2021-11-15 Higher education institutions in Anglophone countries often rely on standardized English language proficiency exams to assess the linguistic capabilities of their multilingual international students However there is often a mismatch between these scores and the initial experiences of international students in both academic and social contexts Drawing on a digital ethnography of Chinese international students first semester languaging practices this book examines their challenges needs and successes on their initial languaging journeys in higher education It analyzes how they use their rich multilingual and multi modal communicative repertoires to facilitate languaging across contexts in order to suggest how university support systems might better serve the needs of multilingual international students

[Chemistry 440 - Sample Midterm Exam](#) Peter Mahaffy,

Student's Guide to Brown and LeMay Chemistry James C. Hill, 1988

McGraw-Hill's 500 Physical Chemistry Questions: Ace Your College Exams Richard H. Langley, 2012-12-12 500 Ways to Achieve Your Best Grades We want you to succeed on your physical chemistry midterm and final exams That s why we ve selected these 500 questions to help you study more effectively use your preparation time wisely and get your best grades These questions and answers are similar to the ones you ll find on a typical college exam so you will know what to expect on test day Each question includes explanations for right and wrong answers for your full understanding of the concepts Whether you have been studying all year or are doing a last minute review McGraw Hill s 500 Physical Chemistry Questions will help you achieve the final grade you desire Sharpen your subject knowledge and build your test taking confidence with 500 essential physical chemistry questions with answers Explanations for every answer Coverage of physical chemistry from ethical theory to epistemology

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Schaum's Outline of Elements of Statistics I: Descriptive Statistics and Probability Stephen Bernstein,Ruth Bernstein,1999-01-21 Tough Test Questions Missed Lectures Not Enough Time Fortunately for you there s Schaum s Outlines More than 40 million students have trusted Schaum s to help them succeed in the classroom and on exams Schaum s is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum s Outline gives you Practice problems with full explanations that reinforce knowledge Coverage of the most up to date developments in your course field In depth review of practices and applications Fully compatible with your classroom text Schaum s highlights all the important facts you need to know Use Schaum s to shorten your study time and get your best test scores Schaum s Outlines Problem Solved *The Writer's World* Linda Woodson,1986 Chemistry James C. Hill,2000 This text defines the concepts needed to learn or review cardiac auscultation The combination of audio and text explains how to identify and interpret normal and common abnormal heart sounds Some heart sounds are reproduced on a heart sound simulator allowing for a clear crisp grasp of specific individual sounds Others are recorded from real patients to distinguish between similar heart and lung sounds and to help the listener select the heart sounds from the auditory milieu

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