

**Question 37** (5 marks)

When performing industrial reductions with  $\text{CO(g)}$ , the following equilibrium is of great importance.



A 1.00 L sealed vessel at a temperature of 1095 K contains  $\text{CO(g)}$  at a concentration of  $1.10 \times 10^{-2} \text{ mol L}^{-1}$ ,  $\text{CO}_2\text{(g)}$  at a concentration of  $1.21 \times 10^{-3} \text{ mol L}^{-1}$ , and excess solid carbon.

- (a) Is the system at equilibrium? Support your answer with calculations.

2

① Calculate Q value

Note: exclude solids from Q expression

$$Q = \frac{[\text{CO}_2]}{[\text{CO}]^2} = \frac{1.21 \times 10^{-3}}{(1.1 \times 10^{-2})^2} = 1.1 \times 10^{-1}$$

② Compare Q to  $K_{eq}$

As  $Q = 1.1 \times 10^{-1} < K_{eq} = 10.00 \therefore$  system is NOT at equilibrium & shifts right.

- (b) Carbon dioxide gas is added to the system above and the mixture comes to equilibrium. The equilibrium concentrations of  $\text{CO(g)}$  and  $\text{CO}_2\text{(g)}$  are equal. Excess solid carbon is present and the temperature remains at 1095 K.

3

Calculate the amount (in mol) of carbon dioxide added to the system.

① use  $K_{eq}$  expression to find equilibrium conc.

$$K_{eq} = \frac{[\text{CO}_2]}{[\text{CO}]^2} = \frac{[\text{CO}_2]}{[\text{CO}_2]^2} = \frac{1}{[\text{CO}_2]} \therefore \frac{1}{[\text{CO}_2]} = 10, [\text{CO}_2] = 0.1$$

② Construct an ICE table ( $[\text{CO}]_{eq} = [\text{CO}_2]_{eq}$ )

	$[\text{CO}]$	$[\text{CO}_2]$
Initial	$1.1 \times 10^{-2}$	$1.21 \times 10^{-3}$
change	+0.0890	+0.0988
Equilibrium	0.1	0.1

③ Complete calculations

change in  $[\text{CO}_2]$  during shift is  $\frac{0.0890}{2} = 0.0445 \text{ mol L}^{-1}$   
 $\therefore$  Amount of  $\text{CO}_2$  to be added =  $0.0988 + 0.0445$  (2:1 ratio)  
 $= 0.143 \text{ mol}$  (1 L container  $\therefore 1 \text{ M} = 1 \text{ mol}$ )

End of paper

# Chemistry Module 12 Answer To Question 9

**Paul Collison, David Kirkby, Averil  
Macdonald**



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