

Question 37 (5 marks)

When performing industrial reductions with CO(g) , the following equilibrium is of great importance.



A 1.00 L sealed vessel at a temperature of 1095 K contains CO(g) at a concentration of $1.10 \times 10^{-2} \text{ mol L}^{-1}$, $\text{CO}_2\text{(g)}$ at a concentration of $1.21 \times 10^{-3} \text{ mol L}^{-1}$, and excess solid carbon.

(a) Is the system at equilibrium? Support your answer with calculations. 2

① Calculate \varnothing value

Note: exclude solids from \varnothing expression

$$\varnothing = \frac{[\text{CO}_2]}{[\text{CO}]} = \frac{1.21 \times 10^{-3}}{1.1 \times 10^{-2}} = 1.1 \times 10^{-1}$$

② Compare \varnothing to K_{eq}

As $\varnothing = 1.1 \times 10^{-1} < K_{eq} = 10.00 \therefore \text{System is NOT at equilibrium & shifts right}$

(b) Carbon dioxide gas is added to the system above and the mixture comes to equilibrium. The equilibrium concentrations of CO(g) and $\text{CO}_2\text{(g)}$ are equal. Excess solid carbon is present and the temperature remains at 1095 K. 3

Calculate the amount (in mol) of carbon dioxide added to the system.

① Use K_{eq} expression to find equilibrium conc.

$$K_{eq} = \frac{[\text{CO}_2]^2}{[\text{CO}]^2} = \frac{[\text{CO}_2]^2}{[\text{CO}_2]^2 \cdot [\text{CO}_2]} \therefore \frac{1}{[\text{CO}_2]} = 10, [\text{CO}_2] = 0.1$$

② Construct an ICE table ($[\text{CO}]_{eq} = [\text{CO}_2]_{eq}$)

	$[\text{CO}]$	$[\text{CO}_2]$
Initial	1.1×10^{-2}	1.21×10^{-3}
change	$+0.0890$	$+0.0988$
Equilibrium	0.1	0.1

③ Complete calculations

Change in $[\text{CO}_2]$ during shift is $\frac{0.0890}{1.1 \times 10^{-2}} = 0.0445 \text{ mol L}^{-1}$

Amount of CO_2 to be added = $0.0988 + 0.0445$ (2:1 ratio)

$$\text{End of paper} = 0.143 \text{ mol} \quad (1 \text{ L container}) \\ \therefore 1 \text{ M} = 1 \text{ mol}$$

Chemistry Module 12 Answer To Question 9

**Paul Collison,David Kirkby,Averil
Macdonald**

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