

- silver nitrate + sodium chloride \longrightarrow silver chloride + sodium nitrate



The spectator ions are NO_3^- and Na^+

(b) Lead (II) iodide can be made by mixing lead (II) nitrate solution with potassium iodide solution.

- Lead (II) nitrate + potassium iodide \longrightarrow Lead (II) iodide + potassium nitrate



- $\text{Pb}(\text{NO}_3)_2_{(\text{aq})} + 2\text{KI}_{(\text{aq})} \longrightarrow \text{PbI}_{2(\text{s})} + 2\text{KNO}_3_{(\text{aq})}$
- The ionic equation is: $\text{Pb}^{2+}_{(\text{aq})} + 2\text{I}^{-}_{(\text{aq})} \longrightarrow \text{PbI}_{2(\text{s})}$
- The spectator ions are NO_3^- and K^+

(c) Calcium carbonate forms on e.g. mixing calcium chloride and sodium carbonate solutions

- Calcium chloride + sodium carbonate \longrightarrow calcium carbonate + sodium chloride



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