

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \qquad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \qquad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \qquad E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left(\begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right) \qquad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \qquad PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \qquad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \qquad P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \qquad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \qquad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} \qquad M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left(\begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left(\begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right) \qquad K_w = [\text{H}^+][\text{OH}^-]$$

$$\left(\begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left(\begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right) \qquad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \qquad \text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left(\begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left(\begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right) \qquad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left(\begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left(\begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right) \qquad \Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

Chemistry Staar Test Review

Reza Nazari, Ava Ross



Chemistry Staar Test Review:

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similar answer choices is more valuable than simply memorizing terms All the material in this book are prepared by chemistry instructors with years of experience in applied chemistry as well as in academic settings This team of experts analyzed the content of the test released by the College Board and designed essential review that will help you build and solidify the knowledge necessary for your success on the exam The content was reviewed for quality and effectiveness by our science editors who possess extensive credentials are educated in top colleges and universities and have years of teaching and editorial experience *Social Sciences & Humanities Index* ,1958 *Sterling Test Prep OAT General Chemistry Review: Complete Subject Review* Sterling Test Prep,2018-01-09 OAT best seller Higher score money back guarantee High yield general chemistry review covering topics tested on OAT Periodic properties and electronic structure Bonding Phases of matter Stoichiometry Thermochemistry and thermodynamics Chemical equilibria and kinetics Solutions Acids and bases Electrochemistry This book provides a detailed and thorough review of general chemistry topics tested on the OAT The content covers foundational principles and theories necessary to answer related questions on the test The information is presented clearly and organized in a systematic way to provide students with targeted OAT review tool You can focus on one knowledge area at a time to learn and fully comprehend important concepts and principles or to simply refresh your memory By reading these review chapters thoroughly you will learn important chemistry concepts and the relationships between them This will prepare you for the OAT and you will significantly increase your score All the material in this book is prepared by chemistry instructors with years of experience This team of experts analyzed the content of OAT test and developed general chemistry review material that will help you build knowledge and develop the skills necessary for your success on the test The material was reviewed and systematized by our OAT editors to ensure strict adherence to the topics and skills outlined by the ADA for the current OAT Our editors are experts on teaching preparing students for standardized tests and have coached thousands of applicants on admission strategies *Abstracts in Biocommerce* ,1989 **Current Index to Journals in Education** ,1974

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