

9. If the pH of a sulfuric acid solution is 1.2, What is the concentration of hydronium?

$$[H_3O^+] = 10^{-1.2} = .063 M$$

What is the hydroxide concentration?

$$[OH^-] = 10^{-12.8} = 1.58 \times 10^{-13} M$$

What is the molarity of the solution?

$$\frac{.063}{2} = .0315 M$$

What is the pOH?

$$14 - 1.2 = 12.8$$

10. If the pH of a solution of sodium hydroxide is 9.0, What is the concentration of hydronium?

$$[H_3O^+] = 10^{-9.0} = 1.0 \times 10^{-9} M$$

What is the hydroxide concentration?

$$[OH^-] = 10^{-5.0} = 1.0 \times 10^{-5} M$$

What is the molarity of the solution?

$$1.0 \times 10^{-5} M$$

What is the pOH?

$$14 - 9.0 = 5.0$$

11. If the pH of a solution of ammonia is 7.6, What is the concentration of hydronium?

$$[H_3O^+] = 10^{-7.6} = 2.5 \times 10^{-8} M$$

What is the hydroxide concentration?

$$[OH^-] = 10^{-6.4} = 4.0 \times 10^{-7} M$$

What is the molarity of the solution?

$$4.0 \times 10^{-7} M$$

What is the pOH?

$$14 - 7.6 = 6.4$$

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John A. Olmsted, Gregory M. Williams



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