

HubSpot

# How To Design a Productivity System That Works for You

Featuring tips and tricks from content creator, designer, and business owner Oliur Online



# Ebook Productivity Guide

**ihabhu**

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2  $\text{CO}_2 \text{ g} \rightarrow \text{H}_2 \text{O g}$   $\Delta h = -283 \text{ kJ}$   
2  $\text{H}_2 \text{O g} \rightarrow \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g}$   $\Delta h = +285 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 3 \text{H}_2 \text{O g}$   $\Delta h = -1560 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$   $\Delta h = +1411 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + \text{O}_2 \text{ g} \rightarrow \text{CO}_2 \text{ g} + \text{H}_2\text{O l}$   $\Delta h = -1560 \text{ kJ}$   
 $\text{H}_2\text{O l} \rightarrow \text{H}_2\text{O g}$   $\Delta h = +285 \text{ kJ}$

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web hess s law states that if a process can be expressed as the sum of two or more steps the enthalpy change for the overall process is the sum of the  $\Delta h$  values for each step to use hess s law two principles must be understood one if an equation is reversed the sign of the  $\Delta h$  value is also reversed two if an equation is multiplied by a

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web thus for reaction 9 2 in which 1 mole of  $H_3O$  and 1 mole of  $OH$  combine to form 2 moles of  $H_2O$   $h 55.8\text{ kJ}$  as shown note that the enthalpy change  $h$  products reactants is positive if heat is absorbed that is if products

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $1 NaOH(aq) + HCl(aq) \rightarrow 1 H_2O(aq)$   $h 2.0\text{ kJ}$   $1 NaOH(aq) + HCl(aq) \rightarrow 1 H_2O(aq)$   $h 2.3\text{ kJ}$

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