

Covalent Bonding

How does a covalent bond form?

- A **covalent bond** forms when two atoms **share one or more pairs of electrons**.
- Atoms in covalent bonds typically have **similar electronegativities** and are **usually non-metals**.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

Types of Covalent Bonds

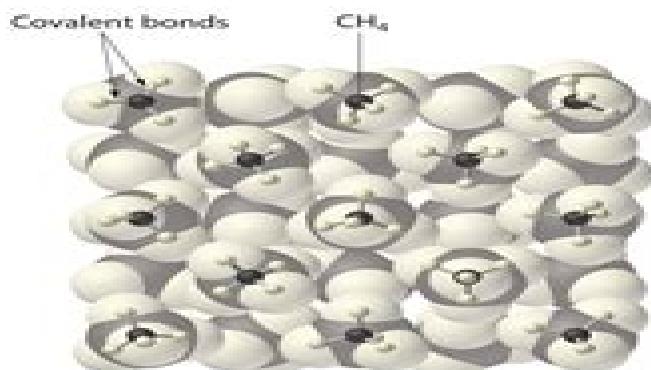
- Single Bond:** One pair of electrons is shared (e.g., H_2 , Cl_2).
- Double Bond:** Two pairs of electrons are shared (e.g., O_2 , CO_2).
- Triple Bond:** Three pairs of electrons are shared (e.g., N_2).



Dot-and-Cross Diagrams

- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

Example: Water H_2O



Properties of Covalent Compounds

- Molecular Structure:** They are made up of discrete molecules with **weak intermolecular forces**.
- Melting and Boiling Points:** Generally **lower** than ionic compounds because the forces between molecules are weaker.
- Electrical Conductivity:** **Poor conductors** of electricity in all states, as they do not have free ions or electrons.
- Solubility:** Often **soluble in nonpolar solvents** and may be insoluble in water, depending on their polarity.

Covalent Bonding Study Guide

Kenneth W. Watkins

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Table of Contents Covalent Bonding Study Guide

1. Understanding the eBook Covalent Bonding Study Guide
 - The Rise of Digital Reading Covalent Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Study Guide
 - Personalized Recommendations
 - Covalent Bonding Study Guide User Reviews and Ratings
 - Covalent Bonding Study Guide and Bestseller Lists
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- Covalent Bonding Study Guide eBook Subscription Services
- Covalent Bonding Study Guide Budget-Friendly Options

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- Covalent Bonding Study Guide Compatibility with Devices
- Covalent Bonding Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Covalent Bonding Study Guide
- Highlighting and Note-Taking Covalent Bonding Study Guide
- Interactive Elements Covalent Bonding Study Guide

8. Staying Engaged with Covalent Bonding Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Covalent Bonding Study Guide

9. Balancing eBooks and Physical Books Covalent Bonding Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Covalent Bonding Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Covalent Bonding Study Guide

- Setting Reading Goals Covalent Bonding Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Covalent Bonding Study Guide

- Fact-Checking eBook Content of Covalent Bonding Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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