



Covalent Compounds Unit Study Guide

Name _____
 Teacher _____
 Date: _____
 Period _____

Reference:

Chemistry in the Community, pp 181-183,
 Notes
 Naming covalent compounds worksheet
 Viscosity Lab
 Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

Covalent	Octet rule
Valence electrons	Polymers
e dot diagrams (Lewis dot diagrams)	Viscosity
Polar	Boiling points
Nonpolar	
Intermolecular forces	

Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
 - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"
 Metal cations & nonmetal anion (or polyatomic ion)
 Net charge = zero (criss cross charges) $X^{+n}Y^{-m} = X_nY_m$
 Dissociate (dissolve) in water
 - > Molecular compounds = sharing electrons by forming a covalent bond
 2 or more nonmetals
 Prefixes indicate # of each atom
 Polar molecules have a δ^+ & a δ^-
 Polar molecules dissolve in water
 Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
 - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

Atomic structure	Solubility
Significant figures	Safety
Accuracy & precision	Elements
Mixtures	Ionic compound formula writing
Nomenclature	

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Arihant Experts



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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011 *Study Guide and Solutions Manual for*

Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006

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Study Guide and Full Solutions Manual Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

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