



Covalent Compounds Unit Study Guide

Name _____
Teacher _____
Date: _____
Period: _____

Reference:

Chemistry in the Community, pp. 181-183.

Notes

Naming covalent compounds worksheet

Viscosity Lab

Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

Covalent

Octet rule

Valence electrons

Polymers

e dot diagrams (Lewis dot diagrams)

Viscosity

Polar

Boiling points

Nonpolar

Intermolecular forces

Targeted Skills

1. Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
2. Know the following properties:
 - Ionic compounds = the transfer of electron(s), i.e. "gain or lose"
Metal cations & nonmetal anion (or polyatomic ion)
Net charge = zero (criss cross charges): $X^{+m}Y^{-n} = X_mY_n$
Dissociate (dissolve) in water
 - Molecular compounds = sharing electrons by forming a covalent bond
2 or more nonmetals
Prefixes indicate # of each atom
Polar molecules have a δ^+ & a δ^-
Polar molecules dissolve in water
Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
3. Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
4. Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
5. Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
6. Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
7. Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

Atomic structure
Significant figures
Accuracy & precision
Mixtures
Nomenclature

Solubility
Safety
Elements
Ionic compound formula writing

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Table of Contents Covalent Compounds Study Guide

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 - The Rise of Digital Reading Covalent Compounds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Compounds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Compounds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Compounds Study Guide

- Personalized Recommendations
- Covalent Compounds Study Guide User Reviews and Ratings
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- Highlighting and Note-Taking Covalent Compounds Study Guide
- Interactive Elements Covalent Compounds Study Guide

8. Staying Engaged with Covalent Compounds Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Covalent Compounds Study Guide

9. Balancing eBooks and Physical Books Covalent Compounds Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Covalent Compounds Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Covalent Compounds Study Guide

- Setting Reading Goals Covalent Compounds Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Covalent Compounds Study Guide

- Fact-Checking eBook Content of Covalent Compounds Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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