



## Reaction Rates and Equilibrium Study Guide

### Topics Covered:

Factors affecting reaction rates  
Collision theory of reaction rates  
Reaction progress diagrams (activation energy and transition states)  
Rate laws and the method of initial rates  
Chemical equilibrium  
Equilibrium expressions and constants and their meaning  
Product-favored vs. Reactant-favored  
Shifting equilibrium and Le Chatelier's Principle  
Quantitative calculations of equilibrium systems using ICE tables

### Practice Questions: [Click here for an answer key](#)

- Describe how each of the following factors would affect the rate of "particle collisions" in your response.
  - Temperature decrease
  - Addition of iodide ions,  $I^-$ , as a catalyst
  - Increased reactant concentration
  - Crushing a solid into a powder (with larger surface area)
- The reaction progress diagram shows the energy changes that take place in the reaction of carbon monoxide gas with nitrogen dioxide gas.



# Chapter Assessment Reaction Rates Study Guide

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## Chapter Assessment Reaction Rates Study Guide:

*Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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