

## Chemistry 162 Study Guide (first exam)

Useful equations:  $S = k_d(P)$ ,  $[A]_t = -kt + [A]_0$ ,  $P_{\text{total}} = X_A(P_A) + X_B(P_B)$ ,  $1/[A]_t = kt + 1/[A]_0$ ,  $\Delta T = K_{\text{fus}}(m)$ ,  
 $q = m(C_p)\Delta T$ ,  $\ln[A]_t = -kt + \ln[A]_0$ ,  $P_{\text{osmosis}} = MRT$ ,  $R = 0.08206 \text{ atm}\cdot\text{L}/\text{mol}\cdot\text{K}$ ,  $T(\text{K}) = T(^{\circ}\text{C}) + 273.15$

1.) Which of the following is a liquid at room temperature (only one is a liquid)?

- a)  $\text{CH}_3\text{OH}$     b)  $\text{CH}_3\text{CH}_3$     c)  $\text{CH}_3\text{CH}_2\text{F}$     d)  $\text{CH}_3\text{F}$     e)  $\text{CH}_3\text{OCH}_3$     f)  $\text{CH}_4$

Intermolecular attractions hydrogen bonding is the strongest force

2.) Viscosity is the measure of the resistance of a liquid to flow, true or false? (Circle one)

No equations or anything just need to know what it is...

3.) What compound will **not** exhibit hydrogen bonding in the liquid state?

- a)  $\text{CH}_3-\text{CH}_2-\text{Br}$     b)  $\text{CH}_3-\text{CH}_2-\text{NH}_2$     c)  $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{OH}$     d)  $\text{CH}_3-\text{NH}-\text{CH}_3$     e)  $\text{NH}_2-\text{O}-\text{H}$

A is the only molecule which doesn't have X-H bond, X = N, O, or F

4.) Which compound cannot be liquefied by compression at temperature of  $25.0^{\circ}\text{C}$ ?

a.  $\text{CH}_3\text{Cl}$ , critical point:  $144.0^{\circ}\text{C}$ , 66.0 atm

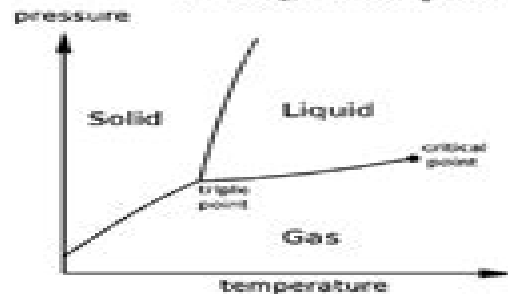
b.  $\text{C}_2\text{H}_2$ , critical point:  $35.5^{\circ}\text{C}$ , 61.6 atm

c.  $\text{C}_2\text{H}_6$ , critical point:  $32.2^{\circ}\text{C}$ , 48.2 atm

d.  $\text{SiF}_4$ , critical point:  $-14.1^{\circ}\text{C}$ , 36.7 atm.  $25^{\circ}\text{C}$  is beyond the critical point thus

only a supercritical fluid is possible... see a phase diagram

e.  $\text{SO}_2$ , critical point:  $158.0^{\circ}\text{C}$ , 78.0 atm



5.) The solubility of  $\text{O}_2$  in water is approximately  $0.00380 \text{ g L}^{-1}$  when the temperature is  $25.0^{\circ}\text{C}$  and the partial pressure of gaseous oxygen is 760 torr. What will the solubility of oxygen be if the oxygen pressure is readjusted to 1000 torr? Henry's law ( $C = kP$ ) but the substance is the same so constant drops out.

a.  $0.00289 \text{ g L}^{-1}$

b.  $0.00500 \text{ g L}^{-1}$  Pg 490  $C_1/P_1 = C_2/P_2$   $C_2 = 0.0038(1000)/760 = 0.0050 \text{ g L}^{-1}$

c.  $1.49 \text{ g L}^{-1}$

d.  $2.89 \times 10^3 \text{ g L}^{-1}$

e.  $3.46 \times 10^3 \text{ g L}^{-1}$

6.) A solution is made by mixing 138.2 grams of ethanol,  $\text{C}_2\text{H}_5\text{O}$ , ( $46.069 \text{ g mol}^{-1}$ ); 103.6 grams of water ( $18.015 \text{ g mol}^{-1}$ ), and 80.11 grams of methanol,  $\text{CH}_3\text{O}$  ( $32.042 \text{ g mol}^{-1}$ ). What is the mole fraction of methanol in the mixture? Calc mol of each 3 mol EtOH, 5.75 water, 2.5 MeOH.

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**G. Marc Loudon, Jim Parise**



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