

## Chemistry 162 Study Guide (first exam)

Useful equations:  $S = k_a(P)$ ,  $[\text{A}]_t = -kt + [\text{A}]_0$ ,  $P_{\text{total}} = X_A(P_A) + X_B(P_B)$ ,  $1/[\text{A}]_t = kt + 1/[\text{A}]_0$ ,  $\Delta T = K_{\text{cool}}(m)$ ,  $q = m(C_s)\Delta T$ ,  $\ln[\text{A}]_t = -kt + \ln[\text{A}]_0$ ,  $P_{\text{partial}} = MRT$ ,  $R = 0.08206 \text{ atm}\cdot\text{L/mol}\cdot\text{K}$ ,  $T(\text{K}) = T(\text{^\circ C}) + 273.15$

1.) Which of the following is a liquid at room temperature (only one is a liquid)?  
 a)  $\text{CH}_3\text{OH}$    b)  $\text{CH}_3\text{CH}_3$    c)  $\text{CH}_3\text{CH}_2\text{F}$    d)  $\text{CH}_3\text{F}$    e)  $\text{CH}_3\text{OCH}_3$    f)  $\text{CH}_4$

Intermolecular attractions hydrogen bonding is the strongest force

2.) Viscosity is the measure of the resistance of a liquid to flow, true or false? (Circle one)  
No equations or anything just need to know what it is...

3.) What compound will not exhibit hydrogen bonding in the liquid state?

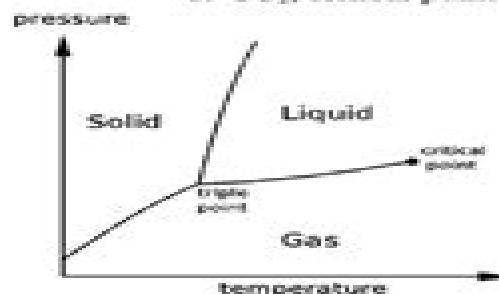
a)  $\text{CH}_3\text{CH}_2\text{Br}$    b)  $\text{CH}_3\text{CH}_2\text{NH}_2$    c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$    d)  $\text{CH}_3\text{NHCH}_3$    e)  $\text{NH}_3\text{O}-\text{H}$

A is the only molecule which doesn't have X-H bond, X = N, O, or F

4.) Which compound cannot be liquefied by compression at temperature of 25.0  $^\circ\text{C}$ ?

a.  $\text{CH}_3\text{Cl}$ , critical point: 144.0  $^\circ\text{C}$ , 66.0 atm  
 b.  $\text{C}_2\text{H}_2$ , critical point: 35.5  $^\circ\text{C}$ , 61.6 atm  
 c.  $\text{C}_2\text{H}_6$ , critical point: 32.2  $^\circ\text{C}$ , 48.2 atm

d.  $\text{SiF}_4$ , critical point: -14.1  $^\circ\text{C}$ , 36.7 atm. 25  $^\circ\text{C}$  is beyond the critical point thus only a supercritical fluid is possible... see a phase diagram  
 e.  $\text{SO}_3$ , critical point: 158.0  $^\circ\text{C}$ , 78.0 atm



5.) The solubility of  $\text{O}_2$  in water is approximately 0.00380 g  $\text{L}^{-1}$  when the temperature is 25.0  $^\circ\text{C}$  and the partial pressure of gaseous oxygen is 760 torr. What will the solubility of oxygen be if the oxygen pressure is readjusted to 1000 torr? Henry's law (C = kP) but the substance is the same so constant drops out.

a. 0.00289 g  $\text{L}^{-1}$   
 b. 0.00500 g  $\text{L}^{-1}$   $P_g = 490 \text{ C}_1/P_1 = C_2/P_2$ ,  $C_2 = 0.0038(1000)/760 = 0.0050 \text{ g L}^{-1}$   
 c. 1.49 g  $\text{L}^{-1}$   
 d.  $2.89 \times 10^3 \text{ g L}^{-1}$   
 e.  $3.46 \times 10^3 \text{ g L}^{-1}$

6.) A solution is made by mixing 138.2 grams of ethanol,  $\text{C}_2\text{H}_5\text{O}$ , (46.069 g  $\text{mol}^{-1}$ ); 103.6 grams of water (18.015 g  $\text{mol}^{-1}$ ), and 80.11 grams of methanol,  $\text{CH}_3\text{O}$  (32.042 g  $\text{mol}^{-1}$ ). What is the mole fraction of methanol in the mixture? Calc mol of each 3 mol EtOH, 5.75 water, 2.5 MeOH.

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**G. Marc Loudon, Jim Parise**

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