



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant, K, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graph showing concentration vs. time for a reaction at equilibrium. The concentration of reactants decreases and the concentration of products increases until they reach a constant value.</p>
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant, K_c, is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant, K_p, is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
<p>Equilibrium Constant, K</p> <p>The equilibrium constant, K, is defined as the ratio of the activities of the products to the activities of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{a_C^c a_D^d}{a_A^a a_B^b}$	<p>Equilibrium Constant, K</p> <p>The equilibrium constant, K, is defined as the ratio of the activities of the products to the activities of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{a_C^c a_D^d}{a_A^a a_B^b}$	
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>		
<p>Concentration</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products.Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 will shift the equilibrium to the right, favoring the formation of NH_3.</p>
<p>Pressure</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the pressure will shift the equilibrium to the side with the fewer moles of gas.Decreasing the pressure will shift the equilibrium to the side with the more moles of gas.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure will shift the equilibrium to the right, favoring the formation of NH_3.</p>
<p>Temperature</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the temperature will shift the equilibrium to the side that is endothermic.Decreasing the temperature will shift the equilibrium to the side that is exothermic.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature will shift the equilibrium to the left, favoring the reactants.</p>

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