

Study Guide- Chemical Reactions Test-December 17, 2014

Chemical Interactions Textbook pages (Chapter 1: Section 1, Section 2, Section 3): 14-21; 24-28; 32-37

Notebook pages:

Review Study guide/ Notes from Quiz 1- Chemical Reactions/Bonding, Logs of scientific terms (Chapter 1 section 1 and section 2)

Review the study guide Monday Dec. 15 :**Review game** on Tuesday December 16th. **Extra help-** during Flex, see me if want to come in during lunch.

You will be able to use your periodic table book mark. You will NOT be permitted to go to your locker if you forget it the day of the test.

Key Terms:

Ionic Bonds	Covalent Bonds	Molecule	Compound
Chemical Equations	Subscript	Coefficient	Reactants
Products	Law of Conservation of Mass	Exothermic Reaction	Endothermic Reaction
Catalyst	Inhibitor	Concentration	Surface Area
Temperature	Synthesis	Decomposition	Replacement

Answer on separate sheet of paper:

1) Define Ionic and Covalent Bonds. What are the differences between them?

Ionic bonds transfer (valance) electrons between a metal and a non-metal during a chemical reaction. A covalent bond is the sharing of (valence) electrons between two non-metals during a chemical reaction.

2) Name the 5 different types of evidence for chemical reactions.

Change in temperature-

Change in properties- light or odor(smell) is given off, change in texture (cake batter (goopy —> fluffy cake)

Change of color

Forms a solid (precipitate) during a chemical reaction

Gas production- which can be seen from fizzing or bubbling

3) Explain the difference between a molecule and a compound.

A molecule is when two or more atoms bond together

A compound is when two or more atoms combine chemically (bond) together in a specific ratio

4) Identify the coefficient and subscript in: 3CHCl_3

Coefficient \rightarrow 3CHCl_3 \leftarrow subscript

5) What information is included in a chemical formula? In a chemical equation?

A chemical formula is a combination of chemical (element) symbols, (elements and subscripts that represent the numbers of atoms in that element),that represent the number of elements in a compound.

Chemical Reactions Test Study Guide Answers

Tyler L. Jackson



Chemical Reactions Test Study Guide Answers:

Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

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