

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

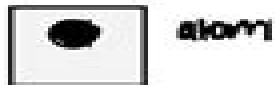
Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
 - o **Daltons Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
 - o **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

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G Orfield

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Excel Science Study Guide, Years 9-10 Will Marchment,2004 The book contains coverage of five major topic areas in the NSW School Certificate test Energy Force and Motion Atoms Elements and Compounds Structure and Function of Living Things Earth and Space Ecosystems Resources and Technology a chapter on Investigations and Problem Solving in Science to help with practical skills revision questions and chapter tests to help you remember important information a glossary and summary in each section of the book diagrams and illustrations to help your understanding a section to help you prepare for the School Certificate test a sample School Certificate test paper with answers answers to all questions *Teaching First Year Chemistry* Joseph Otto Frank,1927

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