

## Chemistry Study Guide: Unit 6: Gases

**Kinetic Molecular Theory (KMT):** a model or theory that is used to explain the behavior of gases.

- A. Gases contain particles (molecules or atoms) that are in constant, random, straight-line motion.
- B. Gas particles collide with each other and with the walls of the container. These collisions are perfectly elastic, meaning kinetic energy (energy an object has because of its motion) is being transferred among the particles but no kinetic energy is lost.
- C. Gas particles are separated by relatively great distances. Because of this, the volume of the particles themselves is negligible (insignificant).
- D. Gas particles do not attract or repel each other. This is because they are far away from each other.

**Gas Pressure:** Force exerted by a gas on the walls of its container.

- Caused by collisions between gas particles and walls of the container.
- Units: kilopascal (kPa), Atmosphere (atm), millimeters mercury (mm Hg)
- $1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$
- Found on table A with temperature on periodic table

**Temperature:** a measure of the average kinetic energy of the particles in a substance.

- As average kinetic energy of particles in a substance increases, temperature of a substance increases.
- In science, Celsius and kelvin are common temperature scales.
- $K = C^{\circ} + 273$ ,  $C^{\circ} = K - 273$
- Temperature is not a form of energy.

Relationship between pressure and volume of a gas:

**Boyle's Law:** Pressure and Volume (Temperature Constant)

If the temperature is constant, then the pressure of a gas increases, as the volume decreases. As the pressure decreases the volume increases.

- Inverse relationship between pressure and volume.

**Equation:**  $P_1 \bullet V_1 = P_2 \bullet V_2$

**Charles Law:** Temperature and Volume (Pressure Constant)

If the pressure is constant, then as the temperature of an enclosed gas increases, the volume increases. As the temperature decreases, the volume decreases.

- Direct relationship between temperature and volume
- Temperature must be converted to Kelvin.

**Equation:**  $V_1 / T_1 = V_2 / T_2$

**Gay-Lussac's Law:** Temperature and Pressure (Volume is constant)

If the volume is constant, then as the temperature of an enclosed gas increases, the pressure increases. As the temperature decreases, the pressure decreases.

- Direct relationship between temperature and pressure.
- Temperature must be converted to Kelvin.

**Equation:**  $P_1 / T_1 = P_2 / T_2$

# Chemistry Gases Unit Study Guide

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## Chemistry Gases Unit Study Guide:

**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Study Guide [to] Chemistry, a Systematic Approach [by] Sisler, Dresdner, and Mooney* William H. Myers, William T. Mooney, 1980 **Resources in education**, 1984-06 *Study Guide* Darrell D. Ebbing, Susan M. Schelble, 1998 **Gage Canadian GED Study Guide Series** Kathy Austin, 1998

Chemistry - Study Guide to Accompany Moore/Davies/Collins Clare T. Furse, 1978 **McGraw-Hill Encyclopedia of Science and Technology** Daniel N. Lapedes, 1977 Accompanied by Readers guide 33 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 1 Accompanied by Study guide 50 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 2 **Gas Chromatography** Ian A. Fowles, 1995-06-14 This volume has been revised and updated to include new chapters on high resolution gas chromatography GC GC detectors and the classification of GC Self assessment questions are used throughout the text to ensure a full understanding of unfamiliar concepts El-Hi Textbooks in Print, 1980 **Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition** Susan McMurry, 2000

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**Chemistry** John S. Phillips, Cheryl Wistrom, 2000 Gas Journal, 1930 **Resources in Education**, 1982 *Medical Technology Examination Review and Study Guide* Anna P. Ciulla, Georganne K. Buescher, Joan H. Youse, 1988 *The Indiana Teacher*, 1936 Iron & Coal Trades Review, 1929 Vol 115 includes Diamond jubilee issue 1867 1927 **Study Guide for Chemical Principles [by]**

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## **Table of Contents Chemistry Gases Unit Study Guide**

1. Understanding the eBook Chemistry Gases Unit Study Guide
  - The Rise of Digital Reading Chemistry Gases Unit Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Gases Unit Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Gases Unit Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Gases Unit Study Guide
  - Personalized Recommendations
  - Chemistry Gases Unit Study Guide User Reviews and Ratings
  - Chemistry Gases Unit Study Guide and Bestseller Lists
5. Accessing Chemistry Gases Unit Study Guide Free and Paid eBooks
  - Chemistry Gases Unit Study Guide Public Domain eBooks
  - Chemistry Gases Unit Study Guide eBook Subscription Services
  - Chemistry Gases Unit Study Guide Budget-Friendly Options
6. Navigating Chemistry Gases Unit Study Guide eBook Formats

- ePub, PDF, MOBI, and More
  - Chemistry Gases Unit Study Guide Compatibility with Devices
  - Chemistry Gases Unit Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Gases Unit Study Guide
    - Highlighting and Note-Taking Chemistry Gases Unit Study Guide
    - Interactive Elements Chemistry Gases Unit Study Guide
  8. Staying Engaged with Chemistry Gases Unit Study Guide
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Gases Unit Study Guide
  9. Balancing eBooks and Physical Books Chemistry Gases Unit Study Guide
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Gases Unit Study Guide
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Gases Unit Study Guide
    - Setting Reading Goals Chemistry Gases Unit Study Guide
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Gases Unit Study Guide
    - Fact-Checking eBook Content of Chemistry Gases Unit Study Guide
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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