

CONCENTRATION OF SOLUTIONS



The concentration of a solution is usually measured in moles per cubic decimetre (mol/dm³). This is a measure of the number of moles in one cubic decimetre.

The volume must be in dm³ (there are 1000 cm³ in 1 dm³). $\text{vol in dm}^3 = \frac{\text{vol in cm}^3}{1000}$

$$\text{concentration (mol/dm}^3) = \frac{\text{moles}}{\text{volume (dm}^3)}$$



1) Calculate the concentration of the following solutions in mol/dm³.

a) 0.10 moles of NaCl in 200 cm³

b) 0.20 moles of H₂SO₄ in 100 cm³

c) 0.500 moles of NaOH in 25 cm³

2) Calculate the number of moles in the following solutions.

a) 100 cm³ of 0.20 mol/dm³ HNO₃

b) 25 cm³ of 1.50 mol/dm³ KOH

c) 50 cm³ of 0.10 mol/dm³ H₂SO₄

Concentration can also be measured in grams per cubic decimetre (g/dm³). This is a measure of the number of grams in one cubic decimetre. (Remember that mass = M = moles)

1 dm³

2 moles of H₂SO₄
196 g. of H₂SO₄

Concentration = 2 mol/dm³

196 g. of H₂SO₄ = 196

Concentration = 2 × 196 = 392 g/dm³

A simple conversion is:

$$\text{conc (g/dm}^3) = \text{conc (mol/dm}^3) \times M$$

3) Calculate the concentration of the following solutions in g/dm³.

a) 0.100 mol/dm³ NaOH

b) 0.250 mol/dm³ CH₃COOH

c) 1.50 mol/dm³ HNO₃

4) 0.20 moles of NaCl is dissolved in 250 cm³ of water.

a) Calculate the concentration in mol/dm³

b) Calculate the concentration in g/dm³

5) 5.0 g of KNO₃ is dissolved in 100 cm³ of water.

a) Calculate the concentration in g/dm³

b) Calculate the concentration in mol/dm³

Chemistry Solutions Study Guide Answers

K. Peter C. Vollhardt, Neil E. Schore

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