

## General Chemistry II Chapter 15 Study Guide

- $S'_{\text{gas}} > S'_{\text{liquid}} > S'_{\text{solid}}$
- For similar molecules,  $S'$  tends to increase with:
  - Increasing molar mass
  - Increased number of atoms in the formula

### Standard entropy of a reaction ( $\Delta S'_{\text{rxn}}$ ):

- $\Delta S'_{\text{rxn}}$ : is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S'_{\text{rxn}} = \Sigma m S'_{\text{products}} - \Sigma n S'_{\text{reactants}}$$

where  $m$  and  $n$  are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction:  $a A + b B \rightarrow c C + d D$

$$\Sigma m S'_{\text{products}} = c \Delta S'_{\text{rxn}} (C) + d \Delta S'_{\text{rxn}} (D)$$

$$\Sigma n S'_{\text{reactants}} = a \Delta S'_{\text{rxn}} (A) + b \Delta S'_{\text{rxn}} (B)$$

Remember to multiply  $\Delta H^\circ$  of each substance by its coefficient.

### Predicting the sign of $\Delta S'_{\text{rxn}}$ :

- We can often predict the sign of  $\Delta S'_{\text{rxn}}$  for processes that involve a change in the number of moles of gas
  - $\Delta S'_{\text{rxn}}$  is positive if the amount of gas increases
  - $\Delta S'_{\text{rxn}}$  is negative if the amount of gas decreases
  - $\Delta S'_{\text{rxn}}$  is likely to be positive if a new structure forms that has more freedom of motion

### Gibbs free energy:

- Gibbs free energy ( $G$ ): combines the enthalpy and entropy of a system:
  - $G = H - TS$
- Gibbs free energy change ( $\Delta G$ ): is a measure of the spontaneity of a process and of the useful energy available from it:
  - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$  for a spontaneous reaction ( $-\Delta G$ )
- $\Delta G > 0$  for a nonspontaneous reaction ( $+\Delta G$ )
- $\Delta G = 0$  for a process at equilibrium ( $\Delta G = 0$ )

# Chemistry Chapter 15 Study Guide

**Margaret G. Kimble**

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