

## General Chemistry II Chapter 15 Study Guide

- $S_{\text{gas}} > S_{\text{liquid}} > S_{\text{solid}}$
- For similar molecules,  $S^\circ$  tends to increase with:
  - Increasing molar mass
  - Increased number of atoms in the formula

Standard entropy of a reaction ( $\Delta S_{\text{rxn}}^\circ$ ):

- $\Delta S_{\text{rxn}}^\circ$ : is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S_{\text{rxn}}^\circ = \sum m S_{\text{products}}^\circ - \sum n S_{\text{reactants}}^\circ$$

where  $m$  and  $n$  are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction:  $a \text{ A} + b \text{ B} \rightarrow c \text{ C} + d \text{ D}$

$$\sum m S_{\text{products}}^\circ = c \Delta S^\circ (\text{C}) + d \Delta S^\circ (\text{D})$$

$$\sum n S_{\text{reactants}}^\circ = a \Delta S^\circ (\text{A}) + b \Delta S^\circ (\text{B})$$

Remember to multiply  $\Delta H^\circ$  of each substance by its coefficient.

Predicting the sign of  $\Delta S_{\text{rxn}}^\circ$ :

- We can often predict the sign of  $\Delta S_{\text{rxn}}^\circ$  for processes that involve a change in the number of moles of gas
  - $\Delta S_{\text{rxn}}^\circ$  is positive if the amount of gas increases
  - $\Delta S_{\text{rxn}}^\circ$  is negative if the amount of gas decreases
  - $\Delta S_{\text{rxn}}^\circ$  is likely to be positive if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy ( $G$ ): combines the enthalpy and entropy of a system:
  - $G = H - TS$
- Gibbs free energy change ( $\Delta G$ ): is a measure of the spontaneity of a process and of the useful energy available from it:
  - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$  for a spontaneous reaction ( $-\Delta G$ )
- $\Delta G > 0$  for a nonspontaneous reaction ( $+\Delta G$ )
- $\Delta G = 0$  for a process at equilibrium ( $\Delta G = 0$ )

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**Margaret G. Kimble**



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