

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent
insoluble	solid	solvent

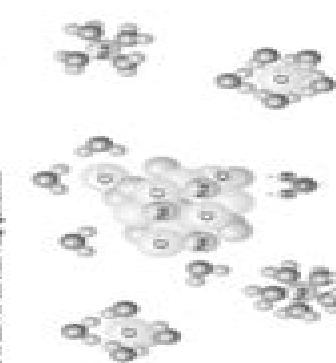
After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is dissolved in the (3) _____ solvent _____.

(4) _____ solute _____ is dissolved in a solvent, or (5) _____ is in (6) _____ solution. A substance that does not dissolve is (7) _____ insoluble _____ in (8) _____ liquid _____ solution or the most common type of solution. If one liquid is soluble in another liquid, such as water and oil, the two liquids are (9) _____ miscible _____ because if one liquid is insoluble in another, the liquids are (10) _____ immiscible _____.

Read about solubility in aqueous solution in your textbook.

Use the diagram to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.

1. Hydration is solvation in which the solute is a solid. What is solvation?
Solution is the process of surrounding solute particles with solvent particles to form a solution.

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Chemistry: Matter and Change • Chapter 15

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CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY**Section 15.1** *Answers*

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecule and the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. In a glass cup, add the rest of solute may be dissolved.

Starting at the bottom of the solution, breaking the solution into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvents, with tips or jobs.

True 16. The overall ionic charge that occurs when a solute dissolves is lost of solutes.

False 17. Solubility is a property of the solute, amount of either the solute or a given amount of solute at a specified temperature and pressure.

False 18. Solvent molecules is large as the solute but is less than the crystal lattice size.

True 19. In a saturated solution, solute and crystallization are in equilibrium.

True 20. Additional solute can be dissolved in an unsaturated solution.

True 21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Hector M. Adams

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Study Guide to DSM-5-TR® Laura Weiss Roberts, Alan K. Louie, Matthew L. Edwards, 2024-12-12 The ultimate companion volume to DSM 5 TR the Study Guide is designed to help clinical learners teachers and practitioners in psychiatry psychology and social work understand and apply diagnostic criteria Medical students residents and fellows and trainees throughout the clinical neurosciences will find this text to be invaluable in developing their diagnostic abilities Readers learn key clinical content through concise chapters and can assess their knowledge with more than 100 multiple choice questions The book is organized into three sections Part I delves into foundational concepts of psychiatric diagnoses and explores how diagnostic assessment fits into the process of working therapeutically with patients In this new edition Part I includes a new chapter examining structural and cultural considerations in arriving at a diagnosis Part II focuses on the diagnostic classes within DSM 5 TR and features clinical case vignettes with real life complexity relevant to psychiatric diagnoses diagnostic pearls that can be applied to daily clinical practice and a self assessment section with key concepts questions for further discussion patient cases and Q As Part III includes more than 100 multiple choice questions with an answer key These questions cover a range of diagnoses and patient presentations from across DSM 5 TR helping readers review their learning and assess their skills Taking a patient centered approach that complements the more disorder centered organization of DSM 5 TR this Study Guide provides context for DSM 5 TR's diagnoses while offering an array of interesting and memorable cases Rich in thought provoking detail and designed for repeated use this is an indispensable guide for learners educators and clinicians

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