

Chemistry: Ionic & Covalent Bonds Unit Study Guide

Covalent Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Covalent Compounds.

- Bond (Ionic / Covalent)
- Charge (+) / (-)
- Atoms (2+) / (2-)
- (High / Low) Conductivity
- (Solid / Liquid / Gas) State
- Polarity (Yes / No)
- Nonpolar

2. Define the following numbers

State	Number
Solid	1
Liquid	2
Gaseous	3
Plasma	4

3. What are the two types of covalent bonds?

4. Complete the following table.

FORMULA	NAME
HF ₂	Hydrogen Fluoride
Br ₂	Chlorine
Si ₂	Silicon
Li	Lithium

5. What does VSEPR stand for?

Name: _____ Date: _____ Period: _____

Ionic Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Ionic Compounds.

- Bond (between a (Metal / Nonmetal) and (Metal / Nonmetal))
- (Large / Small) Electronegativity Difference
- The (Metal / Nonmetal) form Cations, which (Gains / Loses) Electrons
- The (Metal / Nonmetal) form Anions, which (Gains / Loses) Electrons
- (High / Low) Melting & Boiling Point
- (Conducts / Does not Conduct) Electricity in Solution
- (Solid / Liquid / Gas) at Room Temperature

2. What charges would the following periodic table elements have in a chemical bond?

Period	Charge
Metalloids	
Alkali Metals	-1
Alkaline Earth Metals	-2
Halogens	-3
Nitrogen	-4

3. Write the atomic number that is represented by the following numbers.

Atomic Number	Number
1	Hydrogen
2	Helium
3	Lithium
4	Boron
5	Carbon
6	Nitrogen
7	Fluorine
8	Chlorine
9	Silicon
10	Neon

4. Complete the chart below by filling in the missing boxes.

SATION	ANION	FORMULA	NAME
Mg ²⁺	Br ⁻		
		CuO	
			Iron (III) chloride
		(NH ₄) ₂ PO ₄	

Covalent Bonds Study Guide

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**

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Table of Contents Covalent Bonds Study Guide

1. Understanding the eBook Covalent Bonds Study Guide
 - The Rise of Digital Reading Covalent Bonds Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonds Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonds Study Guide
 - Personalized Recommendations
 - Covalent Bonds Study Guide User Reviews and Ratings
 - Covalent Bonds Study Guide and Bestseller Lists

5. Accessing Covalent Bonds Study Guide Free and Paid eBooks
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 - Covalent Bonds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonds Study Guide
 - Highlighting and Note-Taking Covalent Bonds Study Guide
 - Interactive Elements Covalent Bonds Study Guide
8. Staying Engaged with Covalent Bonds Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonds Study Guide
9. Balancing eBooks and Physical Books Covalent Bonds Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonds Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Covalent Bonds Study Guide
 - Setting Reading Goals Covalent Bonds Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Covalent Bonds Study Guide
 - Fact-Checking eBook Content of Covalent Bonds Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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