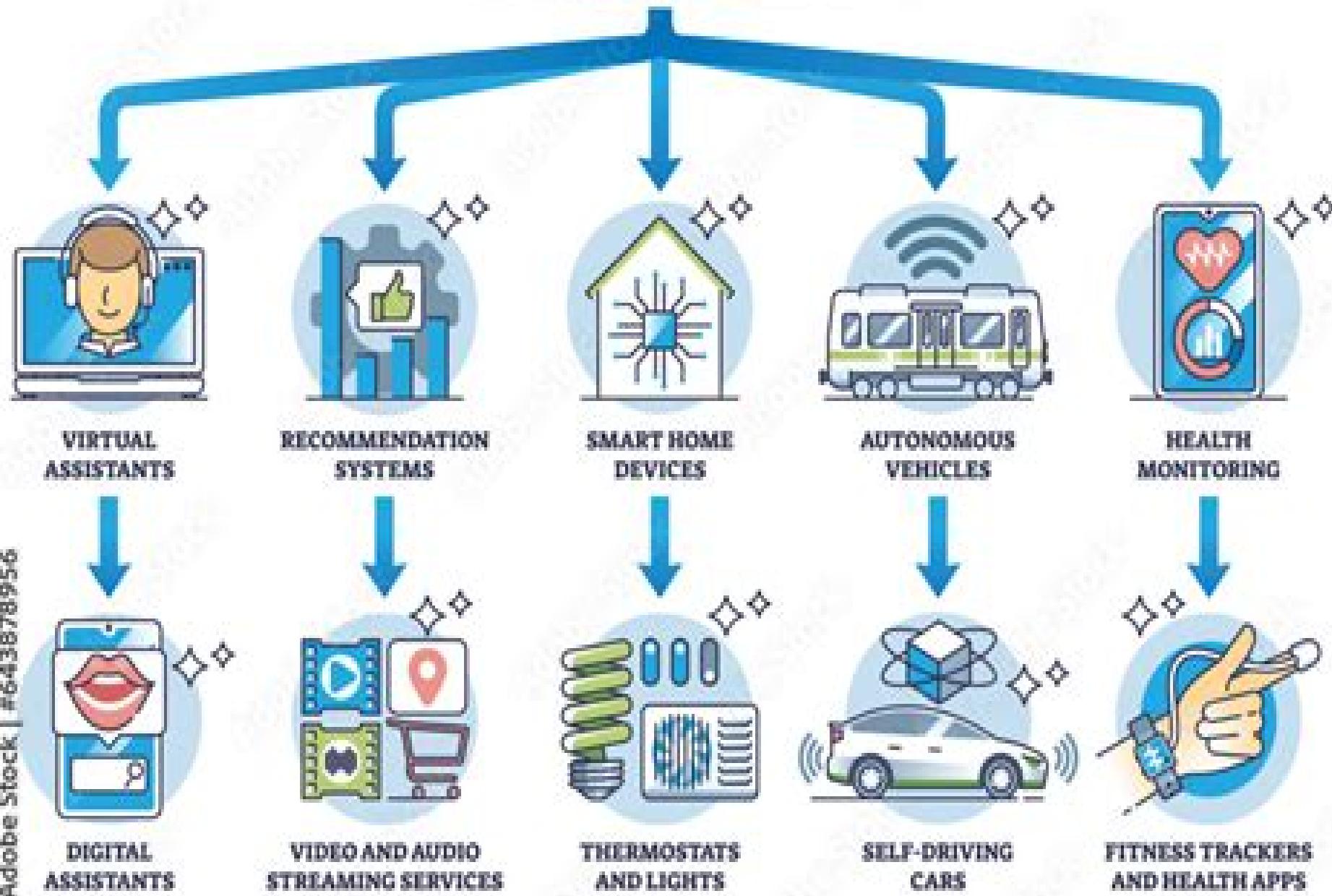


# AI IN DAILY LIFE



# AI In Everyday Life Blueprint

**Richie Cooley**

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the result of this bond is a n bond neutral the sum of the charges on the ions in a unit of the compound is molecule are neutral particles formed as a result of sharing electrons covalent bond a is the force of attraction between atoms sharing electrons bonds

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in the n 2 molecule the nitrogen atoms have an  $\sigma$  bond and two  $\pi$  bonds holding the two atoms together the presence of three strong bonds makes n 2 a very stable molecule phosphorus is a third period element and as such does not form  $\pi$  bonds

efficiently therefore it must fulfill its bonding requirement by forming three  $\sigma$  bonds

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complete atoms can or electrons to get a stable outer energy level gain lose share a figure that uses the symbol of the element and dots to represent the electrons in the outer energy level also used to show how valence electrons are bonded when elements combine to form compounds

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