

CHEMICAL REACTIONS OF COPPER AND PERCENT YIELD

KEY

Objective

To gain familiarity with basic laboratory procedures, some chemistry of a typical transition element, and the concept of percent yield.

Apparatus and Chemicals

0.5 g piece of no. 16 or no. 18 copper wire	evaporating dish
250 mL beaker (2)	weighing paper
concentrated HNO_3 (4 – 6 mL)	6.0 M H_2SO_4 (15 mL)
graduated cylinder	granular zinc
3.0 M NaOH (30 mL)	methanol
carbonundum boiling chips	acetone
stirring rod	towel
iron ring and ring stand	balance
wire gauze	aluminum foil cut in 1-inch squares
Bunsen burner	concentrated HCl (drops)

Discussion

Most chemical synthesis involves separation and purification of the desired product from unwanted side products. Some methods of separation, such as filtration, sedimentation, decantation, extraction, and sublimation were discussed earlier. This experiment is designed as a quantitative evaluation of your individual laboratory skills in carrying out some of these operations. At the same time you will become more acquainted with two fundamental types of chemical reactions – redox reactions and metathesis (double-displacement) reactions. By means of these reactions, you will finally recover the copper sample with maximum efficiency. The chemical reactions involved are the following.

$\text{Cu(s)} + 4 \text{HNO}_3(\text{aq}) \longrightarrow \text{Cu}(\text{NO}_3)_2(\text{aq}) + 2 \text{NO}_2(\text{g}) + 2 \text{H}_2\text{O(l)}$	Redox	[1]
$\text{Cu}(\text{NO}_3)_2(\text{aq}) + 2 \text{NaOH}(\text{aq}) \longrightarrow \text{Cu}(\text{OH})_2(\text{s}) + 2 \text{NaNO}_3(\text{aq})$	Metathesis	[2]
$\text{Cu}(\text{OH})_2(\text{s}) \longrightarrow \text{CuO(s)} + \text{H}_2\text{O(g)}$	Dehydration	[3]
$\text{CuO(s)} + \text{H}_2\text{SO}_4(\text{aq}) \longrightarrow \text{CuSO}_4(\text{aq}) + \text{H}_2\text{O(l)}$	Metathesis	[4]
$\text{CuSO}_4(\text{aq}) + \text{Zn(s)} \longrightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu(s)}$	Redox	[5]

Each of these reactions proceeds to completion. Metathesis reactions proceed to completion whenever one of the components is removed from the solution, such as in the formation of a gas or an insoluble precipitate (driving forces). This is the case for reaction [1], [2], and [3], where in reactions [1] and [3] a gas and in reaction [2] an insoluble precipitate are formed. Reaction [5] proceeds to completion because zinc has a lower ionization energy or oxidation potential than copper.

Copper Yield From Iron Experiment

**California Agricultural Experiment
Station**

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Experiment Station Record United States. Office of Experiment Stations,1910 **The Principles of Chemistry**
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