

Intro. to Chemical Reactions

Electrolyte Reactions

1. Write down formulas for the following compounds:

a) Barium hydroxide $\text{Ba}(\text{OH})_2$

b) Sodium phosphate Na_3PO_4

c) Magnesium phosphate $\text{Mg}_3(\text{PO}_4)_2$

d) Calcium iodide CaI_2

2. Write the formulas for the following acids:

a) HClO₄ Hydrogen perchlorate

b) H₂SO₄ Hydrogen sulfate

c) H₃PO₄ Hydrogen phosphate

d) H₂SiO₃ Hydrogen silicic acid

Electrolyte Reactions: Physical Theory Review

There is often a historical confusion between these terms in the following:



There are two ways you can write the chemical equation above:

Conventional notation results in the products and reactants and coefficients.

There is also an older nomenclature known as oxides and hydrates but these are no longer used.

Conventional Notation: Examples

1. Given the following reaction, write the equation using conventional notation:



Conventional notation results in the products and reactants and coefficients.

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