

Cell Study Guide Answers

1. Cell Membrane- Located as the outer membrane of the animal cell and located under the cell wall in the plant cell. A semipermeable membrane made up of protein and fat that surrounds the cell.

Cell Wall- Located only in plant cells. A thick, rigid membrane that gives the cell most of its support and structure.

Chloroplasts- Located only in plant cells. A disc shaped organelle that contains chlorophyll and manufactures food through photosynthesis.

Chromosomes- Located in animal and plant cells *inside the nucleus*. A threadlike material that contains the DNA and is located inside the nucleus.

Cytoplasm- Located in animal and plant cells. A jellylike material inside the cell where all the organelles are located.

Endoplasmic Reticulum- Located in the animal and plant cells. An organelle that transports materials throughout the cell. The rough ER has ribosomes attached to it, the smooth ER does not.

Golgi Bodies- Located in animal and plant cells. An organelle that packages and moves proteins out of the cell.

Nuclear Membrane- Located in animal and plant cells. A semipermeable membrane that protects and guards the nucleus while keeping the DNA inside it.

Nucleus- Located in animal and plant cells. A spherical organelle that contains other organelles within it, including the nucleolus and chromosomes. It controls many functions of the cell.

Vacuole- Located in the animal and plant cells. The animal cells contain many, while the plant cell only has one central vacuole. A fluid filled membrane that stores food and waste temporarily.

Nucleolus- Located in animal and plant cells *within the nucleus*. An organelle located inside the nucleus that controls the cell's activities, it also codes for the synthesis of rRNA.

Lysosomes- Located in animal and plant cells. A spherical organelle where the digestion of cell nutrients takes place.

2. Cell Theory:

1. All organisms are made of one or more cells.
2. The cell is the basic unit of organization of all organisms.
3. All cells come from other cells already in existence.

3. Robert Hooke- First scientist to name cells. Named after seeing a piece of cork – dead cells

Anton Van Leeuwenhoek- 1675- described the first living cells.

Louis Pasteur- Discovered Pasteurization – kills bacteria

4. Cell- The basic unit of structure and function in living things.

5. Plant Cell- Cell Wall and Chloroplast **Animal Cell-** Cell Membrane and shape

Cells System Study Guide Answers

David A. Stetler

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Study Guide for Psychology David G. Myers, Richard O. Straub, 2009-04-03 Longtime Myers collaborator Richard Straub provides an updated study guide for the new edition **Study Guide for Psychology in Everyday Life** David G. Myers, Richard O. Straub, 2008-11-28 **Phlebotomy Exam Review, Enhanced Edition** Ruth E. McCall, 2020-06-19 Phlebotomy Exam Review Enhanced Seventh Edition thoroughly prepares students for any of the national certification exams in phlebotomy **Brain, Mind, and Behavior Study Guide** Joyce Norman, 2000-09-22 This carefully crafted study guide helps students to read and retain text material and provides them with a multitude of learning tools. The Study Guide includes new labeling exercises of important physiological and neurological structures. Each chapter includes a review of key concepts, guided study questions, practice tests, and section reviews that encourage students' active participation in the learning process **Study Guide** Kelly Bouas-Henry, 2005 By Kelly Bouas Henry Missouri Western State College and Douglas A. Bernstein, the Study Guide employs numerous techniques to help students learn. Each chapter, including the optional Industrial Organizational Psychology chapter, contains a detailed outline, a key terms section that presents fresh examples, and learning aids plus a fill in the blank test learning objectives, a concepts and exercises section that shows students how to apply their knowledge of psychology to everyday issues and concerns, a critical thinking exercise, and personal learning activities. In addition, each chapter concludes with a two part self quiz consisting of 40 multiple choice questions. An answer key tells the student not only which response is correct but also why each of the other choices is wrong and quiz analysis tables enable students to track patterns to their wrong answers either by topic or by type of question definition, comprehension, or application **Psychology, Eighth Edition, in Modules Study Guide** Richard Straub, 2006-08-03 Longtime Myers collaborator Richard Straub's study guide is customized to follow the modular format and contents of the text.

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and

percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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stirring the solution adding water then removing water adding acetone heating the acetone on low

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