

Thermochemistry Study Guide

Qualification: Write the Scientific Notation & use it to answer the questions below

Scientific Notation

1. 100 grams of a substance at 100°C, calculate the specific heat

2. 100g of mercury at 100°C, calculate the specific heat

Reading Comprehension: Label the

Chemical Bonds

Reaction



1. How much energy is released?

2. Calculate the energy change

Calculations: Answer the

1. 100g of water at 100°C, calculate the specific heat

2. 100g of mercury at 100°C, calculate the specific heat

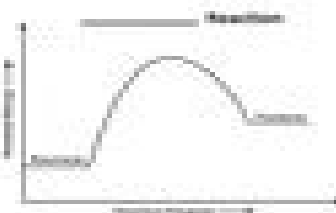
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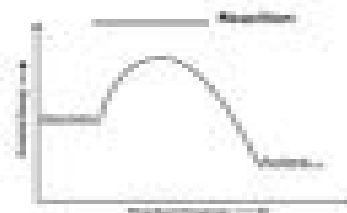
Fill in the blanks with the correct word from the following word bank to complete the statements.
(Word Bank: heat, chemical reactions, light, thermodynamics, energy, temperature, transformation, chemical)

1. According to the law of _____, energy cannot be created or destroyed, only _____ from one form to another.
2. _____ is the ability to do work. _____ energy is stored in chemical bonds. _____ energy is due to the motion of particles and is felt as heat. _____ energy is one we can use and is created by electromagnetic waves.
3. _____ is the measurement of heat energy which can be measured using thermometers while _____ is the energy transferred due to temperature differences.
4. There is thermal _____ when both the system and the surroundings have the same temperature.

Identify which Potential Energy Diagram is Endothermic and which is Exothermic and fill in the blanks to compare their properties.



Heat is _____
_____ Chemical Bonds
Q is _____
Example: _____



Heat is _____
_____ Chemical Bonds
Q is _____
Example: _____

Chemistry Study Guide For Equilibrium And Thermochemistry

Lingsheng Yao



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