

ACS PHYSICAL CHEMISTRY THERMOCHEMISTRY STUDY GUIDE

- Which direction does the reaction proceed if ΔG is positive?: To the left / to reactants
- Which direction does the reaction proceed if ΔG is negative?: To the right/ to products
- How to predict ΔS (system)? (change in entropy of a system): Positive ΔS (system) means available microstates increases/disorder increases
- How to calculate the ΔS° (aka standard entropy change) of the formation of one mole of gas from its elements?: 1. Write the balanced chemical equation with the proper coefficients that make the product ONE MOLE.

2. Find the value of the one mole of product in Appendix B. This is the ΔS° (standard) for the product.

3. Find the value of the reactants in Appendix B for ΔS° (standard) and multiply them by the coefficients you got in step 1.

4. Use total = products-reactants for your final value.

- How to calculate ΔG (free energy) using enthalpy ΔH and entropy ΔS values?:-

$$\Delta G = \Delta H - T\Delta S$$

- What is the equation that relates ΔG to K ? $\Delta G = -RT \ln K$

- What is the Boltzmann constant?: $\ln(\# \text{ of microstates})$

- What does a big W mean?: More microstates

- What does more microstates mean?: Big ΔS - more entropy

- What does fewer microstates mean?: Smaller ΔS - less entropy

- What happens to entropy when volume is increased?: Microstates increase, entropy increases

- What is the relationship between ΔS

of a system, q , and T ? ΔS

$$(\text{system}) = q_{\text{rev}}/T$$

- What happens to entropy ΔS when temperature increases?: It increases

- What is the relationship between mass and entropy ΔS ? As mass increases, entropy increases

- What is the relationship between molecular complexity and entropy ΔS ? As molecular complexity increases, entropy increases

- What is the relationship between ΔG and Q ? $\Delta G = RT \ln Q$

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Williams & Wilkins Review

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