

ACS PHYSICAL CHEMISTRY THERMOCHEMISTRY STUDY GUIDE

- Which direction does the reaction proceed if G is positive?: To the left / to reactants
- Which direction does the reaction proceed if G is negative?: To the right/ to products
- How to predict $S(\text{system})$? (change in entropy of a system): Positive $S(\text{system})$ means available microstates increases/disorder increases
- How to calculate the S° (aka standard entropy change) of the formation of one mole of gas from its elements?: 1. Write the balanced chemical equation with the proper coefficients that make the product ONE MOLE.

2. Find the value of the one mole of product in Appendix B. This is the $\Delta S^\circ(\text{standard})$ for the product.

3. Find the value of the reactants in Appendix B for $\Delta S^\circ(\text{standard})$ and multiply them by the coefficients you got in step 1.

4. Use $\text{total} = \text{products} - \text{reactants}$ for your final value.

- How to calculate G (free energy) using enthalpy H and entropy S values?-

$$\Delta G = \Delta H - T\Delta S$$

- What is the equation that relates G to K ?: $\Delta G = -RT \ln K$
- What is the Boltzmann constant?: $\ln(\# \text{ of microstates})$
- What does a big W mean?: More microstates

- What does more microstates mean?: Big S° - more entropy

- What does fewer microstates mean?: Smaller S° - less entropy

- What happens to entropy when volume is increased?: Microstates increase, entropy increases

- What is the relationship between S

of a system, q , and T ?: S

$$(\text{system}) = q_{\text{rev}}/T$$

- What happens to entropy S when temperature increases?: It increases

- What is the relationship between mass and entropy S ?: As mass increases, entropy increases

- What is the relationship between molecular complexity and entropy S ?: As molecular complexity increases, entropy increases

- What is the relationship between G and Q ?: $\Delta G = RT \ln Q$

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