

# Unit 1 Chemistry Study Guide

## History of the Atom:

### Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

### Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

### Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

### Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

### John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
  - o **Daltons Atomic Theory:**
    - o All matter is composed of extremely small particles called atoms.
    - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
    - o Atoms cannot be subdivided, created, or destroyed.
    - o Atoms to different elements can combine in simple whole number ratios to form chemical compounds.
    - o In chemical reactions, atoms are combined, separated or rearranged.
  - o **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

### Billiard Ball Model:



### J.J. Thompson (1906):

- **Cathode Ray Experiment –**
  - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

# Chemistry Unit Study Guide Answers

**BM King**

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## **Table of Contents Chemistry Unit Study Guide Answers**

1. Understanding the eBook Chemistry Unit Study Guide Answers
  - The Rise of Digital Reading Chemistry Unit Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit Study Guide Answers
  - Personalized Recommendations
  - Chemistry Unit Study Guide Answers User Reviews and Ratings
  - Chemistry Unit Study Guide Answers and Bestseller Lists
5. Accessing Chemistry Unit Study Guide Answers Free and Paid eBooks
  - Chemistry Unit Study Guide Answers Public Domain eBooks
  - Chemistry Unit Study Guide Answers eBook Subscription Services
  - Chemistry Unit Study Guide Answers Budget-Friendly Options

6. Navigating Chemistry Unit Study Guide Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - Chemistry Unit Study Guide Answers Compatibility with Devices
  - Chemistry Unit Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Unit Study Guide Answers
  - Highlighting and Note-Taking Chemistry Unit Study Guide Answers
  - Interactive Elements Chemistry Unit Study Guide Answers
8. Staying Engaged with Chemistry Unit Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Unit Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Unit Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Unit Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit Study Guide Answers
  - Setting Reading Goals Chemistry Unit Study Guide Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit Study Guide Answers
  - Fact-Checking eBook Content of Chemistry Unit Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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