

9. $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$
- At 25°C , 0.11 mole of N_2O_4 reacts to form 0.10 mol of N_2O_4 and 0.02 mole of NO_2 . At 90°C , 0.11 mole of N_2O_4 forms 0.050 mole of N_2O_4 and 0.12 mole of NO_2 . From these data we can conclude
- N_2O_4 molecules react by a second order rate law.
 - N_2O_4 molecules react by a first order rate law.
 - the reaction is exothermic.
 - N_2O_4 molecules react faster at 25°C than at 90°C .
 - the equilibrium constant for the reaction above increases with an increase in temperature.

10. For the equilibrium system



$$\Delta H = -42 \text{ kJ/mol}$$

K_c equals 0.62 at 1260 K. If 0.10 mole each of H_2O , CO , H_2 and CO_2 (each at 1260 K) were placed in a 1.0-Liter flask at 1260 K, when the system came to equilibrium...

	The temperature would	The mass of CO would
a)	decrease	increase
b)	decrease	decrease
c)	remain constant	increase
d)	increase	decrease
e)	increase	increase

$$Q = \frac{(0.10)(0.10)}{(0.10)(0.10)} = 1 \quad K_c = 0.62$$

$Q > K_c$ will shift to
 To use up products
 and make reactants
 Temp $\downarrow$ CO $\uparrow$

11. For the reaction system,



the conditions that would favor maximum conversion of the reactants to products would be

- high temperature and high pressure
- high temperature, pressure unimportant
- high temperature and low pressure
- low temperature and high pressure
- low temperature and low pressure

12. Solid HgO , liquid Hg , and gaseous O_2 are placed in a glass bulb and are allowed to reach equilibrium at a given temperature.



The mass of HgO in the bulb could be increased by

- adding more Hg . NO CHANGE
- removing some O_2 .
- reducing the volume of the bulb.
- increasing the temperature.
- removing some Hg . NO CHANGE

Answers: (Please use CAPITAL letters)

1.	E
2.	D
3.	C
4.	B
5.	D
6.	D

7.	D
8.	A
9.	E
10.	A
11.	D
12.	C

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Table of Contents Chemquest 35 Intro To Gases Answers

1. Understanding the eBook Chemquest 35 Intro To Gases Answers
 - The Rise of Digital Reading Chemquest 35 Intro To Gases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 35 Intro To Gases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 35 Intro To Gases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 35 Intro To Gases Answers
 - Personalized Recommendations

- Chemquest 35 Intro To Gases Answers User Reviews and Ratings
- Chemquest 35 Intro To Gases Answers and Bestseller Lists
- 5. Accessing Chemquest 35 Intro To Gases Answers Free and Paid eBooks
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 - Chemquest 35 Intro To Gases Answers eBook Subscription Services
 - Chemquest 35 Intro To Gases Answers Budget-Friendly Options
- 6. Navigating Chemquest 35 Intro To Gases Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemquest 35 Intro To Gases Answers Compatibility with Devices
 - Chemquest 35 Intro To Gases Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemquest 35 Intro To Gases Answers
 - Interactive Elements Chemquest 35 Intro To Gases Answers
- 8. Staying Engaged with Chemquest 35 Intro To Gases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 35 Intro To Gases Answers
- 9. Balancing eBooks and Physical Books Chemquest 35 Intro To Gases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 35 Intro To Gases Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemquest 35 Intro To Gases Answers
 - Setting Reading Goals Chemquest 35 Intro To Gases Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemquest 35 Intro To Gases Answers
 - Fact-Checking eBook Content of Chemquest 35 Intro To Gases Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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