

Chemistry — Unit 8 Test

1. The following graph shows the relationship between the number of atoms and the mass of an element.

- a. The two axes for this graph are:
- b. The following statement is true concerning this graph:
 - i. The positive slope means increasing the mass does not increase the number of atoms.
 - ii. The negative slope means increasing the mass does not increase the number of atoms.
 - iii. The horizontal slope means the number of atoms does not change.
 - iv. The vertical slope means the number of atoms does not change.

2. The following diagrams show three different ways to represent the same element.

- a. The following statement is true concerning the three diagrams:
 - i. The first diagram is a bar graph.
 - ii. The second diagram is a line graph.
 - iii. The third diagram is a pie chart.
 - iv. The first diagram is a pie chart.
- b. The following statement is true concerning the three diagrams:
 - i. The first diagram is a bar graph.
 - ii. The second diagram is a line graph.
 - iii. The third diagram is a pie chart.
 - iv. The first diagram is a pie chart.
- c. The following statement is true concerning the three diagrams:
 - i. The first diagram is a bar graph.
 - ii. The second diagram is a line graph.
 - iii. The third diagram is a pie chart.
 - iv. The first diagram is a pie chart.



Relationships between atomic and molecular mass

- a. The following statement about atomic mass is true. Explain your answer in terms of the following:
 - i. The mass of an atom.
 - ii. The mass of an element.
 - iii. The mass of a molecule.
 - iv. The mass of a compound.

The following statement about atomic mass is true. Explain your answer in terms of the following:

- i. The mass of an atom.
- ii. The mass of an element.
- iii. The mass of a molecule.
- iv. The mass of a compound.

- b. Formulas containing the symbol H_2O_2 are often used. Explain why this formula is often used. Explain why this formula is often used.

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Annelies Wilder-Smith

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