

## Section 7.6: Solubility Equilibria and the Solubility Product Constant

### Tutorial 1 Practice, page 462

1. **Solution:**  $[\text{MgCO}_3(\text{s})]$  is constant, so it is not included in the equation. Therefore,  $K_{\text{sp}} = [\text{Mg}^{2+}(\text{aq})][\text{CO}_3^{2-}(\text{aq})]$

**Statement:** The solubility product constant equation for the solubility equilibrium is  $K_{\text{sp}} = [\text{Mg}^{2+}(\text{aq})][\text{CO}_3^{2-}(\text{aq})]$

2. **Given:** iron(III) hydroxide dissolves in water

**Required:**  $K_{\text{sp}}$  equation

**Solution:**

Step 1. Write a balanced equation for the dissolution of iron(III) hydroxide in water.



Step 2.  $[\text{Fe(OH)}_3(\text{s})]$  is constant, so it is not included in the equation. Therefore,



**Statement:** The solubility product constant equation for iron(III) hydroxide is



3. **Given:** At 25 °C:  $[\text{Ca}^{2+}(\text{aq})] = 4.53 \times 10^{-7}$  mol/L;  $[\text{PO}_4^{3-}(\text{aq})] = 3.02 \times 10^{-7}$  mol/L

**Required:**  $K_{\text{sp}}$  of calcium phosphate at 25 °C

**Solution:**

Step 1. Write a balanced equation for the solution equilibrium.



Step 2. Write the solubility constant equation for the equilibrium, omitting any solids.



Step 3. Substitute the given ion concentrations into the  $K_{\text{sp}}$  equation and solve.

$$\begin{aligned} K_{\text{sp}} &= [\text{Ca}^{2+}(\text{aq})]^3[\text{PO}_4^{3-}(\text{aq})]^2 \\ &= (4.53 \times 10^{-7})^3(3.02 \times 10^{-7})^2 \end{aligned}$$

$$K_{\text{sp}} = 8.48 \times 10^{-33}$$

**Statement:** The  $K_{\text{sp}}$  of calcium phosphate in a saturated solution at 25 °C is  $8.48 \times 10^{-33}$ .

### Tutorial 2 Practice, page 464

1. **Given:** solubility of  $\text{MgF}_2(\text{s}) = 1.72 \times 10^{-3}$  g/100 mL at 25 °C

**Required:**  $K_{\text{sp}}$  of  $\text{MgF}_2(\text{s})$  at 25 °C

**Solution:**

Step 1. Write the balanced equation.



Step 2. Write the solubility product constant equation.



Step 3. Determine the equilibrium concentrations in mol/L of the dissolved ions. Use an ICE table. Since the concentration of solid  $\text{MgF}_2$  remains constant, it is not included.

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**John Adolphus Flemer**

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