

# Intro. to Chemical Reactions

## Electrolyte Reactions

1. Write down formulas for the following compounds:

a) Magnesium carbonate MgCO<sub>3</sub>

b) Calcium phosphate Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

c) Copper(II) phosphate Cu<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

d) Tin(IV) iodide TiI<sub>4</sub>

2. Write the formulas for the following acids:

a) H<sub>2</sub>SO<sub>4</sub> Hydrogen sulfate

b) H<sub>3</sub>PO<sub>4</sub> Hydrogen phosphate

c) H<sub>2</sub>CO<sub>3</sub> Hydrogen carbonate

d) H<sub>2</sub>SiO<sub>3</sub> Hydrogen silicate

## Electrolyte Reactions: Physical Theory Review

There is often a chemical reaction that does not go forward:



There are reasons that you can name as reasons why things don't go:

**Electrolyte reactions** result when reactants and products are insoluble and unreactive.

There are also no electron relationships between the reactants and the products but this is not always a necessary condition.

## Corrosion: Oxidizing Compounds

1. Given the following reaction, write the equation that goes along with it:



Corrosion products are not always soluble and acidic oxidation reactions are often observed.

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