

# balancing Equations

Name: \_\_\_\_\_

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## Introduction: Balancing

- Write the equation for the reaction of sodium and chlorine (chlorine gas) to form sodium chloride.



- Write the equation for the reaction of sodium oxide and carbon dioxide to produce sodium carbonate sodium.



## Information: Subscripts and Coefficients

It is important to understand that atoms are never created or destroyed in a chemical reaction. For example, in the equation above, if the subscript under NaCl is 1, then there must be 1 sodium atom, 1 chlorine atom, and 1 oxygen atom.

A coefficient tells how many molecules of a substance are present. For example, in the equation, "2 NaCl" the number 2 is the coefficient. The three tells us that there are three molecules of sodium present. In this equation, "1 NaCl" there are a total of 1 sodium atoms and 1 chlorine atoms.

## Balancing Chemical Equations

- How many oxygen atoms are present in  $2Na_2O$ ? 3/1

- How many oxygen atoms are in each of the following:



## Information: How To Balance Equations

Consider the reaction of sodium and chlorine to produce sodium chloride (halogenation):



Remember that when chlorine is the reactant it is always written as  $Cl_2$ . On the product side of the equation, there must always be a total of two atoms of chlorine. Notice the product side there is only one atom of chlorine. Atomic masses are always conserved. In order to balance equation for mass reason, we need to balance the equation. How can we do this? Well, by adding a "2" to the product side:



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