

**Study Guide – Chapter 3: Stoichiometry**

Test Date: Thursday, 12/18/14

- **Chemical Equations**
  - Writing word equations
    - remember to balance & include the states of matter
  - Balancing chemical equations w/coefficients
  - 5 types of reactions
    - Combination/synthesis, Decomposition, Combustion, Single Replacement, Double Replacement
    - identify what type of reaction is happening from a given equation
    - predict the products when given the reactants or told the type of reaction...
      - $\text{C}_2\text{H}_6 + \text{O}_2 \text{ (g)}$  combustion forms products of  $\text{CO}_2 \text{ (g)}$  and  $\text{H}_2\text{O} \text{ (g)}$
      - Synthesis of metal + nonmetal forms an ionic compound (criss-cross charges for formula)
      - Decomposition breaks compound into its elements (watch out for diatomic BrINClHOFs or  $\text{P}_2$  /  $\text{Se}$  solids)
- **Formula Weight (unit: amu) = Molar Mass (unit: g/mol)**
  - Calculations of formula weights and/or molar masses
  - % by mass = % composition calculations
- **Determining Empirical & Molecular Formulas**
  - Know the difference between an empirical and a molecular formula
  - Three types of problems
    - Normal problems will give % composition data or mass data to start
    - Combustion Analysis (like the quiz)
    - Hydrate Determination
      - Use the rhyme to help you remember the steps!
- **Stoichiometric Calculations**
  - Dimensional Analysis!!!
  - Avogadro's number (use for converting to or from # of atoms/molecules) → after molecules you can go to specific atom within the compound by looking at the chemical formula (ex: 4 H atoms per 1 molec.  $\text{CH}_4$ )
  - use the molar mass to go from mass to moles and/or moles to mass
  - Molar Ratios: use coefficients from a balanced chemical equation to go between any two different substances in a reaction (section 3.6)
- **Limiting Reactant Problems**
  - through calculations, identify which is the limiting reactant and which is the excess reactant
  - calculate the amount of product formed by the limiting reactant (this is the maximum or theoretical yield but you **ALWAYS** choose the **LESSER** quantity calculated)
  - determine how much of the excess reactant is used and how much is left over after the reaction is completed
  - calculate % yield
    - calculate theoretical yield and given the actual yield (from lab) of the substance
    - given % yield, you can determine actual yield collected in the lab

# Chapter 11 Stoichiometry Study Guide Answer Key

**A Loxley**

## **Chapter 11 Stoichiometry Study Guide Answer Key:**

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