

Chapter 10: Gases

Learning Outcomes:

- > Calculate pressure and convert between pressure units with an emphasis on torr and atmospheres.
- > Calculate P , V , n , or T using the ideal-gas equation.
- > Explain how the gas laws relate to the ideal-gas equation and apply the gas laws in calculations.
- > Calculate the density or molecular weight of a gas.
- > Calculate the volume of gas consumed or formed in a chemical reaction.
- > Calculate the total pressure of a gas mixture given its partial pressures or given information for calculating partial pressures.
- > Describe the kinetic-molecular theory of gases and how it explains the pressure and temperature of a gas, the gas laws, and the rates of effusion and diffusion.
- > Explain why intermolecular attractions and molecular volumes cause real gases to deviate from ideal behavior at high pressure or low temperature.

1

Characteristics of Gases

Unlike liquids and solids, gases:

- expand spontaneously to fill their containers
- are highly compressible
- have extremely low densities
- have indefinite shape
- can diffuse and mix rapidly with other gases in the same container (different gases in a mixture do not separate upon standing)



2

Chapter 20 Gases Exercises Answers

Joan Bouillon



Chapter 20 Gases Exercises Answers:

Fundamentals of Physics, Chapters 1 - 21 David Halliday, Robert Resnick, Jearl Walker, 2000-05-02 This is a supplement to the text Fundamentals of Physics 6th Ed This supplement contains additional sample problems checkpoint style questions organizing questions discussion questions and new exercises and problems **Fundamentals of Physics, A Student's Companion E-Book to Accompany Fundamentals of Physics, Enhanced Problems Version** David Halliday, Robert Resnick, Jearl Walker, 2003 A text for calculus based physics courses introducing fundamental physics concepts and featuring exercises designed to help students apply conceptual understanding to quantitative problem solving with chapter puzzlers checkpoints and reviews and summaries **Handbook of Chemical Mass Transport in the Environment** Louis J. Thibodeaux, Donald Mackay, 2010-10-21 A comprehensive account of the state of the science of environmental mass transport Edited by Louis J Thibodeaux and Donald Mackay renowned experts in this field the Handbook of Chemical Mass Transport in the Environment covers those processes which are critically important for assessing chemical fate exposure and risk In a comprehensive and authoritative format this unique handbook provides environmental chemists geoscientists engineers and modelers with the essential capabilities to understand and quantify transport In addition it offers a one stop resource on environmental mass transfer and mass transport coefficient estimation methods for all genres The book begins by discussing mass transport fundamentals from an environmental perspective It introduces the concept of mobility key to environmental fate since transport must occur prior to any reaction or partitioning within the natural multimedia compartments The fugacity approach to environmental mass transfer and the conventional approach are examined This is followed by a description of the individual mass transport processes and the appropriate flux equations required for a quantitative expression The editors have identified 41 individual processes believed to be the most environmentally significant which form the basis for the remainder of the book Using a consistent format for easy reference each chapter Introduces the specific processes Provides a detailed qualitative description Presents key theoretical mathematical formulations Describes field or laboratory measurements of transport parameters Gives data tables and algorithms for numerical estimates Offers a guide for users familiar with the process who are seeking a direct pathway to obtain the numerical coefficients Presents computed example problems case studies and or exercises with worked through solutions and answers The final chapter presents the editors insight into future needs and emerging priorities Accessible and relevant to a broad range of science and engineering users this volume captures the state of the transport science and practice in this critical area **Study Guide for Physics in the Modern World 2E** Jerry Marion, 2012-12-02 Study Guide for Physics in the Modern World 2E provides information pertinent to the fundamental concepts in physics This book presents a list of concepts definitions and equations with various supplementary exercises for the readers Comprised of 21 chapters this book starts with an overview of the standard units of measure for length time mass energy force pressure and density This

text then provides the meaning of various terms in physics including atom molecule element and compound Other chapters explore the composition and behavior of all ordinary matter in which it depends on the four basic units including electrons protons neutrons and photons This book discusses as well the method used for converting the units of physical quantities from one system of measurement to another The final chapter deals with the various applications of radiation in biological investigations as well as in medical diagnostics and therapeutics This book is intended for students enrolled in introductory physics courses

Fundamentals of Physics, Part 3, Chapters 22 - 33, Enhanced Problems Version David Halliday, Robert Resnick, 2002-04-16 The primary goal of this text is to provide students with a solid understanding of fundamental physics concepts and to help them apply this conceptual understanding to quantitative problem solving

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Fundamentals of Physics, Part 1, Chapters 1 - 12 David Halliday, Robert Resnick, 2000-01-21 The primary goal of this text is to provide students with a solid understanding of fundamental physics concepts and to help them apply this conceptual understanding to quantitative problem solving

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Physics, Volume 1 Robert Resnick, David Halliday, Kenneth S. Krane, 2002 Each chapter includes questions and problems

Chemistry Stanley R. Radel, Marjorie H. Navidi, 1990 **General College Chemistry** Charles William Keenan, Donald C. Kleinfelter, Jesse Hermon Wood, 1980

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Russian Journal of Physical Chemistry, 1975 New Technical Books New York Public Library, 1925

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