

## General Chemistry II Chapter 13 Study Guide

- Raoult's Law:  $P_a = X_a \cdot P_a^*$ 
  - $P_a$  = vapor pressure of solvent "a" above a solution
  - $P_a^*$  = vapor pressure of pure solvent "a"
  - $X_a$  = mole fraction of solvent "a" in the solution
    - $X_a$  = moles of solvent "a" / moles of solution (solvent particles + solute particles)
- Volatile: substances that easily evaporate and exhibit a vapor pressure
- Nonvolatile: substances that have little tendency to evaporate
- Vapor pressure of the solution is the sum of the partial pressures of the solute and the solvent
  - If solute is non-volatile, vapor pressure of solution only comes from the solvent ( $P_{\text{solution}} < P_{\text{pure solvent}}$ )
  - If solute is more volatile than solvent,  $P_{\text{solution}} > P_{\text{pure solvent}}$

### Review on boiling point:

- Boiling point: temperature where vapor pressure = atmospheric pressure
  - Vapor pressure of liquid increases as temperature increases
  - Starts boiling process when vapor pressure = atmospheric pressure
  - Solute in solution lowers the vapor pressure of solvent- vapor pressure lowering
  - A higher temperature needed for vapor pressure to hit atmospheric pressure
  - boiling point of solution is higher than boiling point of pure solvent- boiling point elevation

### Boiling point elevation:

- boiling point elevation: the boiling point of a solution is always higher than that of the pure solvent; it depends on the type of solvent and the # of solute particles
- $\Delta T_b = iK_b m$ ;  $\Delta T_b = T_b - T_b^*$ 
  - $m$  = solute molality (moles solute / kg solvent)
  - $K_b$  = molal boiling point elevation constant ( $^{\circ}\text{C} / m$ ); depends only on solvent
  - $i$  = van't Hoff factor
  - $T_b$  = boiling point of solution
  - $T_b^*$  = boiling point of pure solvent

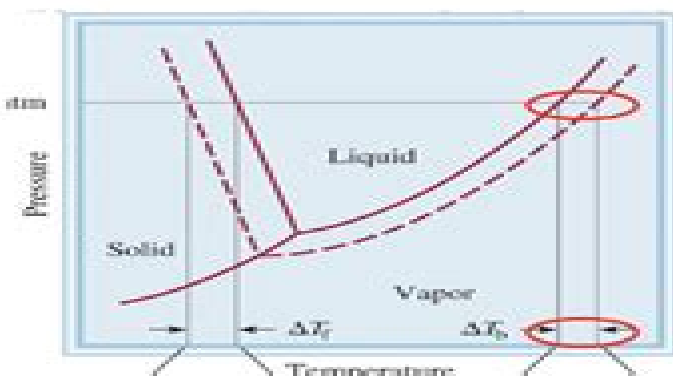


TABLE 13.2

Molal Boiling-Point Elevation and Freezing-Point Depression Constants of Several Common Liquids

Solvent	Normal Freezing Point ( $^{\circ}\text{C}$ )*	$K_f$ ( $^{\circ}\text{C}/m$ )	Normal Boiling Point ( $^{\circ}\text{C}$ )*	$K_b$ ( $^{\circ}\text{C}/m$ )
Water	0	1.86	100	0.52
Benzene	5.5	5.12	80.1	2.53
Ethanol	-117.3	1.99	78.4	1.22
Acetic acid	16.6	3.90	117.9	2.93
Cyclohexane	6.6	20.0	80.7	2.79

# Chapter 13 Study Guide Chemistry Answers

**Richard Watts**



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Ralph J. Fessenden,Joan S. Fessenden,1986 Provides a set of additional drill problems chapter by chapter discussions and supplemental instructional material to help students master organic chemistry problem solving techniques

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## **Table of Contents Chapter 13 Study Guide Chemistry Answers**

1. Understanding the eBook Chapter 13 Study Guide Chemistry Answers
  - The Rise of Digital Reading Chapter 13 Study Guide Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 13 Study Guide Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 13 Study Guide Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 13 Study Guide Chemistry Answers
  - Personalized Recommendations
  - Chapter 13 Study Guide Chemistry Answers User Reviews and Ratings
  - Chapter 13 Study Guide Chemistry Answers and Bestseller Lists
5. Accessing Chapter 13 Study Guide Chemistry Answers Free and Paid eBooks
  - Chapter 13 Study Guide Chemistry Answers Public Domain eBooks
  - Chapter 13 Study Guide Chemistry Answers eBook Subscription Services
  - Chapter 13 Study Guide Chemistry Answers Budget-Friendly Options

6. Navigating Chapter 13 Study Guide Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 13 Study Guide Chemistry Answers Compatibility with Devices
  - Chapter 13 Study Guide Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 13 Study Guide Chemistry Answers
  - Highlighting and Note-Taking Chapter 13 Study Guide Chemistry Answers
  - Interactive Elements Chapter 13 Study Guide Chemistry Answers
8. Staying Engaged with Chapter 13 Study Guide Chemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 13 Study Guide Chemistry Answers
9. Balancing eBooks and Physical Books Chapter 13 Study Guide Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 13 Study Guide Chemistry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 13 Study Guide Chemistry Answers
  - Setting Reading Goals Chapter 13 Study Guide Chemistry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 13 Study Guide Chemistry Answers
  - Fact-Checking eBook Content of Chapter 13 Study Guide Chemistry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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