

## General Chemistry II Chapter 13 Study Guide

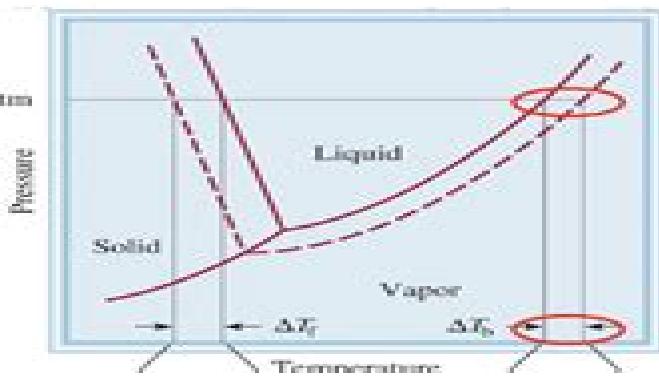
- Raoult's Law:  $P_{\text{sol}} = X_{\text{sol}} \cdot P_{\text{sol}^*}$ 
  - $P_{\text{sol}}$  = vapor pressure of solvent "a" above a solution
  - $P_{\text{sol}^*}$  = vapor pressure of pure solvent "a"
  - $X_{\text{sol}}$  = mole fraction of solvent "a" in the solution
    - $X_{\text{sol}} = \text{moles of solvent "a"} / \text{moles of solution (solvent particles + solute particles)}$
- Volatile: substances that easily evaporate and exhibit a vapor pressure
- Non-volatile: substances that have little tendency to evaporate
- Vapor pressure of the solution is the sum of the partial pressures of the solute and the solvent
  - If solute is non-volatile, vapor pressure of solution only comes from the solvent ( $P_{\text{solution}} < P_{\text{pure solvent}}$ )
  - If solute is more volatile than solvent,  $P_{\text{solution}} > P_{\text{pure solvent}}$

### Review on boiling point:

- Boiling point: temperature where vapor pressure = atmospheric pressure
  - Vapor pressure of liquid increases as temperature increases
  - Starts boiling process when vapor pressure = atmospheric pressure
  - Solute in solution lowers the vapor pressure of solvent- vapor pressure lowering
  - A higher temperature needed for vapor pressure to hit atmospheric pressure
  - boiling point of solution is higher than boiling point of pure solvent- boiling point elevation

### Boiling point elevation:

- boiling point elevation: the boiling point of a solution is always higher than that of the pure solvent; it depends on the type of solvent and the # of solute particles
- $\Delta T_b = iK_b m$  ;  $\Delta T_b = T_b - T_b^*$ 
  - m= solute molality (moles solute / kg solvent)
  - $K_b$ = molal boiling point elevation constant ( $^{\circ}\text{C} / \text{m}$ ); depends only on solvent
  - i= van't Hoff factor
  - $T_b$ = boiling point of solution
  - $T_b^*$ = boiling point of pure solvent



**TABLE 13.2** Molal Boiling-Point Elevation and Freezing-Point Depression Constants of Several Common Liquids

| Solvent     | Normal Freezing Point ( $^{\circ}\text{C}$ )* | $K_f$ ( $^{\circ}\text{C}/\text{m}$ ) | Normal Boiling Point ( $^{\circ}\text{C}$ )* | $K_b$ ( $^{\circ}\text{C}/\text{m}$ ) |
|-------------|---|---------------------------------------|--|---------------------------------------|
| Water       | 0   | 1.86                                  | 100  | 0.52                                  |
| Benzene     | 5.5   | 5.12                                  | 80.1   | 2.53                                  |
| Ethanol     | -117.3  | 1.99                                  | 78.4   | 1.22                                  |
| Acetic acid | 16.6  | 3.90                                  | 117.9  | 2.93                                  |
| Cyclohexane | 6.6   | 20.0                                  | 80.7   | 2.79                                  |

# Chapter 13 Study Guide Chemistry Answers

**Richard Watts**

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