

CHAPTER 16 **STUDY GUIDE****Reaction Rates****Section 16.1 A Model for Reaction Rates**

In your textbook, read about expressing reaction rates and explaining reactions and their rates.

Use each of the terms below just once to complete the passage.

activated complex	activation energy	mol/(L·s)
collision theory	reaction rate	

According to the (1) _____, atoms, ions, and molecules must collide in order to react. Once formed, the (2) _____ is a temporary, unstable arrangement of atoms that may then form products or may break apart to reform the reactants. Every chemical reaction requires energy, and the minimum amount of energy that reacting particles must have to form the activated complex is the (3) _____. In a chemical reaction, the (4) _____ is the change in concentration of a reactant or product per unit time. It may be expressed using the units of (5) _____.

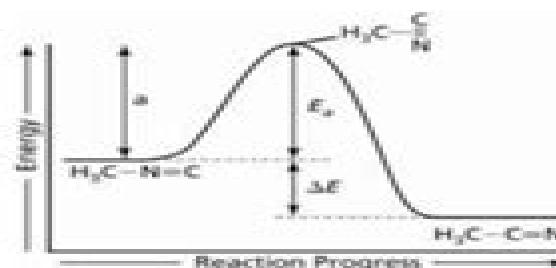
Use the energy diagrams for the rearrangement reaction of methyl isonitrile to acetonitrile to answer the following questions.

6. What kind of reaction is represented by this diagram, endothermic or exothermic?

7. What is the chemical structure identified at the top of the curve on the diagram?

8. What does the symbol E_a represent?

9. What does the symbol ΔE represent?



Chapter 16 Chemistry Study Guide Answers

Bernice G. Segal, Peter S. Shenkin

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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step by step examples. Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Giorgi, 2011-01-12 Chemistry, Study Guide Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry, chemical equilibrium, and thermochemistry. Worked examples now number over 300 and exercises over 1460.

Student Study Guide Fundamentals of Chemistry by David E. Goldberg, Margaret G. Kimble, 1994

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Provides answers and explanations to all in text and end of chapter exercises. Also includes summaries of name reactions, functional group synthesis and reactions, lists of reagents and abbreviations, and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry. This edition now includes all new artwork expanded in text problems, summary quizzes approximately every three chapters, more detailed explanations in solutions and chapter outlines.

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