

# CHAPTER 16 STUDY GUIDE

## Reaction Rates

### Section 16.1 A Model for Reaction Rates

In your textbook, read about expressing reaction rates and explaining reactions and their rates.

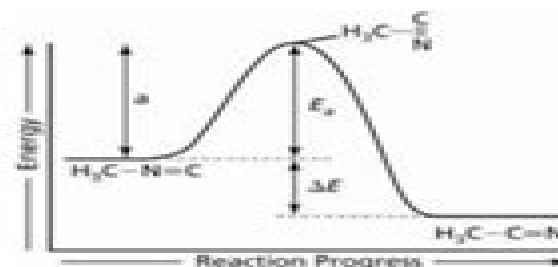
Use each of the terms below just once to complete the passage.

activated complex	activation energy	mol/(L·s)
collision theory	reaction rate	

According to the (1) \_\_\_\_\_, atoms, ions, and molecules must collide in order to react. Once formed, the (2) \_\_\_\_\_ is a temporary, unstable arrangement of atoms that may then form products or may break apart to reform the reactants. Every chemical reaction requires energy, and the minimum amount of energy that reacting particles must have to form the activated complex is the (3) \_\_\_\_\_. In a chemical reaction, the (4) \_\_\_\_\_ is the change in concentration of a reactant or product per unit time. It may be expressed using the units of (5) \_\_\_\_\_.

Use the energy diagram for the rearrangement reaction of methyl isonitrile to acetonitrile to answer the following questions.

- What kind of reaction is represented by this diagram, endothermic or exothermic?  
\_\_\_\_\_
- What is the chemical structure identified at the top of the curve on the diagram?  
\_\_\_\_\_
- What does the symbol  $E_a$  represent?  
\_\_\_\_\_
- What does the symbol  $\Delta E$  represent?  
\_\_\_\_\_



# Chapter 16 Chemistry Study Guide Answers

**Bernice G. Segal, Peter S. Shenkin**



## Chapter 16 Chemistry Study Guide Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Wastewater Operator Certification Study Guide** John

Giorgi, 2011-01-12 Chemistry, Student Study Guide Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first-year chemistry text, known for its clarity of exposition and its large number of illustrative worked problems, contains a more rigorous treatment of electrochemistry, chemical equilibrium, and thermochemistry. Worked examples now number over 300 and exercises over 1460. *Student Study Guide Fundamentals of Chemistry* by David E. Goldberg Margaret G. Kimble, 1994

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**Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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