

Chapter 5: Periodic Law

- I. Worksheet on History of the Periodic Table and Periodic Law.
 - A. Dobereiner, Newlands, Mendeleev, Moseley.
 - B. **Periodic Law:** The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- II. The Periodic Table
 - A. **Periodic Table:** an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column or group.
 - B. Information in each block of the table:
 1. Symbol
 2. Atomic Number
 3. Atomic Mass
 4. Name
 5. Electron Configuration
 - C. The Periodic Table is organized into rows and columns of elements.
 1. **Periods:** Horizontal rows of elements. Numbered 1-7. Those numbers correspond to the principle energy level (n).
 2. **Groups:** Vertical columns of elements. Also called families. numbered in 3 ways.
 - a. **European System:** 1A-VIIIA; IB-VIII B
 - b. **U.S. System:** 1A, 2A, 3B-8B, 1B, 2B, 3A-8A. We will use this one.
 - c. **IUPAC System:** 1-18
 3. Family Names to know:

<u>Group</u>	<u>Family Name</u>
1A	Alkali Metals
2A	Alkaline Earth Metals
7A	Halogens
8A	Noble Gases

Other groups with A are called by the name of the first element in that group. For example, Group 4A is called the Carbon Group (Family).
 4. The elements with B in the group number are the transition metals.
 5. The Lanthanides and Actinides are on the bottom of the table. They are called the inner transition metals. They don't fit into the table very well. They make it too big.

Chapter 6 Study Guide The Periodic Table And Periodic Law Answers

Peter C. Scott



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Table of Contents Chapter 6 Study Guide The Periodic Table And Periodic Law Answers

1. Understanding the eBook Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - The Rise of Digital Reading Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Personalized Recommendations
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers User Reviews and Ratings
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers and Bestseller Lists
5. Accessing Chapter 6 Study Guide The Periodic Table And Periodic Law Answers Free and Paid eBooks
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers Public Domain eBooks
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers eBook Subscription Services
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers Budget-Friendly Options

6. Navigating Chapter 6 Study Guide The Periodic Table And Periodic Law Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers Compatibility with Devices
 - Chapter 6 Study Guide The Periodic Table And Periodic Law Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Highlighting and Note-Taking Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Interactive Elements Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
8. Staying Engaged with Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
9. Balancing eBooks and Physical Books Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Setting Reading Goals Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Fact-Checking eBook Content of Chapter 6 Study Guide The Periodic Table And Periodic Law Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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