

4. a. 
$$\frac{10 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33 \text{ g Cu}$$

$$\frac{20 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45 \text{ g Cu}}$$

b. 
$$\frac{30 \text{ g Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25 \text{ g AlCl}_3$$

$$\frac{100 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11 \text{ g AlCl}_3}$$

c. 
$$\frac{5 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67 \text{ g Cu}$$

~~100 g / 17.67 g~~  $10 \text{ g} / 17.67 \text{ g} \times 100 = \boxed{56.6\%}$

d. 
$$\frac{10 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61 \text{ g AlCl}_3$$

~~100 g / 6.61 g~~  $5 \text{ g} / 6.61 \text{ g} \times 100 = \boxed{75.6\%}$

- 5.
- a. TRUE
  - b. FALSE
  - c.  $\text{AgNO}_3 \Rightarrow 169.91 \text{ g/mol}$
  - d.  $6.02 \times 10^{23}$  particles = 1 mole of particles

# Chapter 11 Study Guide Chemistry Stoichiometry

## Answer Key

**Donald J. Wink, Sharon Fetzer-  
Gislason, Sheila McNicholas**



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### **Table of Contents Chapter 11 Study Guide Chemistry Stoichiometry Answer Key**

1. Understanding the eBook Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - The Rise of Digital Reading Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Study Guide Chemistry Stoichiometry Answer Key

- Personalized Recommendations
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key User Reviews and Ratings
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key and Bestseller Lists
5. Accessing Chapter 11 Study Guide Chemistry Stoichiometry Answer Key Free and Paid eBooks
- Chapter 11 Study Guide Chemistry Stoichiometry Answer Key Public Domain eBooks
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key eBook Subscription Services
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key Budget-Friendly Options
6. Navigating Chapter 11 Study Guide Chemistry Stoichiometry Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key Compatibility with Devices
  - Chapter 11 Study Guide Chemistry Stoichiometry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - Highlighting and Note-Taking Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - Interactive Elements Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
8. Staying Engaged with Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
9. Balancing eBooks and Physical Books Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
- Setting Reading Goals Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 11 Study Guide Chemistry Stoichiometry Answer Key



- Fact-Checking eBook Content of Chapter 11 Study Guide Chemistry Stoichiometry Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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